

## Electron Configuration Practice Worksheet **KEY**

*In the space below, write the unabbreviated electron configurations of the following elements:*

- 1) sodium       $1s^2 2s^2 2p^6 3s^1$
- 2) iron         $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
- 3) bromine     $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$
- 4) barium      $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2$
- 5) neptunium  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 5f^4$

*In the space below, write the abbreviated electron configurations of the following elements:*

- 6) cobalt       $[Ar] 4s^2 3d^7$
- 7) silver       $[Kr] 5s^2 4d^9$
- 8) tellurium  $[Kr] 5s^2 4d^{10} 5p^4$
- 9) radium      $[Rn] 7s^2$
- 10) lawrencium  $[Rn] 7s^2 5f^{14} 6d^1$

Determine what elements are denoted by the following electron configurations:

- 11)  $1s^2 2s^2 2p^6 3s^2 3p^4$  **Sulfur**
- 12)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$  **Rubidium**
- 13)  $[Kr] 5s^2 4d^{10} 5p^3$  **Antimony**
- 14)  $[Xe] 6s^2 4f^{14} 5d^8$  **Osmium**
- 15)  $[Rn] 7s^2 5f^{11}$  **Einsteinium**

Determine which of the following electron configurations are not valid:

- 16)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^5$       **should be  $3d^{10}$**
- 17)  $1s^2 2s^2 2p^6 3s^2 3d^5$                                **$3s^2$  is impossible & "p" instead of "d"**
- 18)  $[Ra] 7s^2 5f^8$                                       **Ra is not a noble gas**
- 19)  $[Kr] 5s^2 4d^{10} 5p^5$                               **Valid**
- 20)  $[Xe]$                                                 **Element can't be its own electron configuration.**

# Electron Configuration Level One Answers

**WJ Hussar**



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## **Table of Contents Electron Configuration Level One Answers**

1. Understanding the eBook Electron Configuration Level One Answers
  - The Rise of Digital Reading Electron Configuration Level One Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Electron Configuration Level One Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Electron Configuration Level One Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Electron Configuration Level One Answers
  - Personalized Recommendations
  - Electron Configuration Level One Answers User Reviews and Ratings
  - Electron Configuration Level One Answers and Bestseller Lists
5. Accessing Electron Configuration Level One Answers Free and Paid eBooks
  - Electron Configuration Level One Answers Public Domain eBooks
  - Electron Configuration Level One Answers eBook Subscription Services
  - Electron Configuration Level One Answers Budget-Friendly Options
6. Navigating Electron Configuration Level One Answers eBook Formats

- ePub, PDF, MOBI, and More
- Electron Configuration Level One Answers Compatibility with Devices
- Electron Configuration Level One Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Electron Configuration Level One Answers
  - Highlighting and Note-Taking Electron Configuration Level One Answers
  - Interactive Elements Electron Configuration Level One Answers
- 8. Staying Engaged with Electron Configuration Level One Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Electron Configuration Level One Answers
- 9. Balancing eBooks and Physical Books Electron Configuration Level One Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Electron Configuration Level One Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Electron Configuration Level One Answers
  - Setting Reading Goals Electron Configuration Level One Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Electron Configuration Level One Answers
  - Fact-Checking eBook Content of Electron Configuration Level One Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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