

Answer Key

PG-1: What is the coefficient of the hydroxide ion for the half reaction when the disproportionation reaction is given the smallest coefficient reduction?  $\text{BrO}^-(\text{aq}) \rightarrow \text{Br}^-(\text{aq}) + \text{BrO}_3^-(\text{aq})$  unbalanced



PG-2: What is the balanced half reaction of the species in reaction in basic solution according to the reaction equation?



PG-3: What is the coefficient for water when the reaction is balanced in acidic solution?  $\text{CrO}_4^{2-}(\text{aq}) + \text{Mn}^{2+}(\text{aq}) \rightarrow \text{Cr}^{3+}(\text{aq}) + \text{MnO}_4^-(\text{aq})$  unbalanced



PG-4: Which reaction occurs in the cell in the presence of standard state conditions?



The standard reduction potential of  $\text{Fe}^{3+}/\text{Fe}^{2+}$  is +0.77 V.  
The standard reduction potential of  $\text{Fe}^{2+}/\text{Fe}$  is 0 V.



The cell is a galvanic cell with  $\text{Fe}^{3+}$  as the oxidizing agent and  $\text{Fe}^{2+}$  as the reducing agent.

Reduction occurs at the cathode of the cell.

For a reaction, the standard cell potential is  $E^\circ_{\text{cell}} = E^\circ_{\text{cathode}} - E^\circ_{\text{anode}}$ .

The standard cell potential is  $E^\circ_{\text{cell}} = 0.77 \text{ V}$ .

# Electrochemistry Answers Study Guide

**TD Snyder**



## **Electrochemistry Answers Study Guide:**

## Embracing the Melody of Expression: An Mental Symphony within **Electrochemistry Answers Study Guide**

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