

**CHEM 113 Final – Monday May 16<sup>th</sup> 10:15-12:15 in LSC 133; bring a pencil!**

**Atomic Structure:**

1. What do these have in common?



- A) The same number of protons
- B) The same number of electrons
- C) The same number of neutrons
- D) The same size

2. The element X occurs natural to the extent of 20.0%  $^{12}\text{X}$  and 80.0%  $^{13}\text{X}$ . The atomic mass of X is nearest

- A) 12.2
- B) 12.5
- C) 12.8
- D) 13.0

3. What is the valence electron configuration for the element in Period 5, Group 3A?

- A)  $5s^25p^1$
- B)  $3s^23p^3$
- C)  $3s^23p^5$
- D)  $5s^25p^3$

4. What is the maximum number of electrons that can occupy an orbital labeled  $d_{xy}$ ?

- A) 1
- B) 2
- C) 4
- D) 10

5. Which electron transition in a hydrogen atom is associated with the largest emission of energy?

- A)  $n = 2$  to  $n = 1$
- B)  $n = 1$  to  $n = 2$
- C)  $n = 2$  to  $n = 4$
- D)  $n = 4$  to  $n = 2$

**Molecular Structure and Bonding:**

1. Which molecule has exactly two unshared (lone) pairs of electrons on the central atom?

- A)  $\text{BF}_3$
- B)  $\text{OF}_2$
- C)  $\text{NF}_3$
- D)  $\text{XeF}_2$

2. The molecule of the type  $\text{ML}_4$  consists of four single bonds and no lone pairs. What structure is it expected to assume?

- A) square planar
- B) trigonal pyramidal
- C) trigonal planar
- D) tetrahedral

# Mastering Chemistry Chem 113 Exercise Answers

**Michael Brown**



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## **Table of Contents Mastering Chemistry Chem 113 Exercise Answers**

1. Understanding the eBook Mastering Chemistry Chem 113 Exercise Answers
  - The Rise of Digital Reading Mastering Chemistry Chem 113 Exercise Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Mastering Chemistry Chem 113 Exercise Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Mastering Chemistry Chem 113 Exercise Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Mastering Chemistry Chem 113 Exercise Answers
  - Personalized Recommendations
  - Mastering Chemistry Chem 113 Exercise Answers User Reviews and Ratings
  - Mastering Chemistry Chem 113 Exercise Answers and Bestseller Lists
5. Accessing Mastering Chemistry Chem 113 Exercise Answers Free and Paid eBooks
  - Mastering Chemistry Chem 113 Exercise Answers Public Domain eBooks
  - Mastering Chemistry Chem 113 Exercise Answers eBook Subscription Services
  - Mastering Chemistry Chem 113 Exercise Answers Budget-Friendly Options
6. Navigating Mastering Chemistry Chem 113 Exercise Answers eBook Formats

- ePub, PDF, MOBI, and More
  - Mastering Chemistry Chem 113 Exercise Answers Compatibility with Devices
  - Mastering Chemistry Chem 113 Exercise Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Mastering Chemistry Chem 113 Exercise Answers
    - Highlighting and Note-Taking Mastering Chemistry Chem 113 Exercise Answers
    - Interactive Elements Mastering Chemistry Chem 113 Exercise Answers
  8. Staying Engaged with Mastering Chemistry Chem 113 Exercise Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Mastering Chemistry Chem 113 Exercise Answers
  9. Balancing eBooks and Physical Books Mastering Chemistry Chem 113 Exercise Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Mastering Chemistry Chem 113 Exercise Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Mastering Chemistry Chem 113 Exercise Answers
    - Setting Reading Goals Mastering Chemistry Chem 113 Exercise Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Mastering Chemistry Chem 113 Exercise Answers
    - Fact-Checking eBook Content of Mastering Chemistry Chem 113 Exercise Answers
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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