



2. What mass in grams of P_4O_6 could be made from 8.75 g P_4 and 12.50 g O_2 ?



$$8.75\text{g P}_4 \times \frac{1\text{mol P}_4}{123.88\text{g}} \times \frac{1\text{mol P}_4\text{O}_6}{1\text{mol P}_4} \times \frac{219.88\text{g}}{1\text{mol P}_4\text{O}_6} = \boxed{15.5\text{g P}_4\text{O}_6}$$

$$12.50\text{g O}_2 \times \frac{1\text{mol O}_2}{32.00\text{g}} \times \frac{1\text{mol P}_4\text{O}_6}{3\text{mol O}_2} \times \frac{219.88\text{g}}{1\text{mol P}_4\text{O}_6} = 28.63\text{g P}_4\text{O}_6$$

Limiting Reagent Question And Answer Key

KJ Lindholm-Leary



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