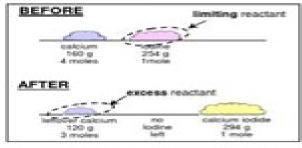
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When an automobile production plant runs out of tires no more cars can be produced even if there are still plenty of engines, bodies, seats, and other parts. When one of the reactants in a chemical reaction runs out the reaction stops as well. The substance that limits production in a chemical reaction is known as the **limiting** reactant. The reactant that is left over when the reaction stops is called the excess reactant.

1 Ca + 1 I₂ → 1 CaI₂

Consider the chemical reaction between calcium and iodine. When 160 g (4 moles) of calcium reacts with 254 g (1 mole) of iodine the reaction makes 294 g (1 mole) of calcium iodide. Although there is enough calcium to make more calcium iodide, the iodine runs out first. Since all of the iodine gets used up it is called the limiting reactant. According to the balanced equation, if one mole of iodine reacts, one mole of calcium will react. This means that there are still 3 moles of calcium left. Because calcium is left over it is called the excess reactant.



Example

What mass of iron (II) sulfide will be produced if 9.68 g of iron reacts with 6.28 g of sulfur? __Fe + __S → __FeS

- balance the equation <u>1</u> Fe + <u>1</u> S → <u>1</u> FeS
- perform a calculation for each reactant: $p: \frac{9.68 \text{ g/Gc}}{1} \times \frac{1 \text{ m/h/fc}}{55.85 \text{ g/Gc}} \times \frac{1 \text{ m/h/fc}}{1 \text{ m/h/fc}} \times \frac{87.91 \text{ g/Fc}}{1 \text{ m/h/fc}} = 15.24 \text{ g/Fc}$
 - S: 6.28 g/s 1 miles 1 mol Fes | 87.91 g FeS | 17.22 g FeS | 1 mol Fes | 1 mol Fes |
- whichever reactant makes less product is
 the limiting reactant. There is enough sulfur to make 17.22 g
 feS, but only enough iron to make 15.24 g of FeS.

Find the amount of product formed in each of the following reactions.

- What mass of H₂O will be produced if 9.5 g of H₂ reacts with 1.2 g of O₂?
 ___ H₂ + ___ O₂ → ___ H₂O
- If 1.85 g of Mg(OH)₂ reacts with 3.71 g of HCl, how much MgCl₂ is produced? What is the limiting reactant? __Mg(OH)₂ + __HCl → __MgCl₂ + __H₂O
- What mass of AgCl is produced when 53.42 g of AgNO₃ reacts with 14.19 g of NaCl? __AgNO₃ + __NaCl → __AgCl + __NaNO₃
- If 14.7 g of calcium is placed in 11.5 g of water, what mass of hydrogen gas is produced? __Ca + __H₂O → __Ca(OH)₂ + __H₂
- What mass of potassium hydroxide is formed when 8.2 g of potassium oxide is added to 1.3 g of water? ___ K₂O + __ H₂O → __ KOH

- What mass of aluminum chloride could be made from 8.1 g of aluminum and 4.2 L of chlorine at STP? __ Al + __ Cl₂ → __ AlCl₃
- If 5.26 L of nitrogen monoxide and 7.64 L of oxygen are combined, what mass of nitrogen dioxide is formed? Assume conditions are STP.
 NO + __O₂ → __NO₂
- If 18.1 g of silicon tetrachloride reacts with 8.4 L of hydrogen at STP, what mass of silicon is formed? __SiCl₄ + __H₂ → __Si + __HCl
- What mass of Cu(OH)₂ is produced in the reaction of 7.6 g Cu(NO₃)₂ with 6.2 g of KOH? _ Cu(NO₃)₂ + _ KOH → _ Cu(OH)₂ + _ KNO₃
- What mass of carbon dioxide is formed when 64 kg of ethylene is burned in 142 kg of oxygen?
 _ C₂H₄ + _ O₂ → _ CO₂ + _ H₂O

Limiting Reactants Chem Worksheet 12 3 Answers

Robert C. Smoot, Smoot, Richard G. Smith, Jack Price

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- 1. Understanding the eBook Limiting Reactants Chem Worksheet 12 3 Answers
 - The Rise of Digital Reading Limiting Reactants Chem Worksheet 12 3 Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Limiting Reactants Chem Worksheet 12 3 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Limiting Reactants Chem Worksheet 12 3 Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Limiting Reactants Chem Worksheet 12 3 Answers
 - Personalized Recommendations
 - Limiting Reactants Chem Worksheet 12 3 Answers User Reviews and Ratings
 - Limiting Reactants Chem Worksheet 12 3 Answers and Bestseller Lists
- 5. Accessing Limiting Reactants Chem Worksheet 12 3 Answers Free and Paid eBooks

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- Limiting Reactants Chem Worksheet 12 3 Answers Budget-Friendly Options
- 6. Navigating Limiting Reactants Chem Worksheet 12 3 Answers eBook Formats
 - o ePub, PDF, MOBI, and More
 - Limiting Reactants Chem Worksheet 12 3 Answers Compatibility with Devices
 - Limiting Reactants Chem Worksheet 12 3 Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Limiting Reactants Chem Worksheet 12 3 Answers
 - Highlighting and Note-Taking Limiting Reactants Chem Worksheet 12 3 Answers
 - Interactive Elements Limiting Reactants Chem Worksheet 12 3 Answers
- 8. Staying Engaged with Limiting Reactants Chem Worksheet 12 3 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Limiting Reactants Chem Worksheet 12 3 Answers
- 9. Balancing eBooks and Physical Books Limiting Reactants Chem Worksheet 12 3 Answers
 - Benefits of a Digital Library
 - o Creating a Diverse Reading Collection Limiting Reactants Chem Worksheet 12 3 Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Limiting Reactants Chem Worksheet 12 3 Answers
 - Setting Reading Goals Limiting Reactants Chem Worksheet 12 3 Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Limiting Reactants Chem Worksheet 12 3 Answers
 - Fact-Checking eBook Content of Limiting Reactants Chem Worksheet 12 3 Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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