

Calculations need to be based on the limiting reactant.

- Example 2: What is the mass of sulfur dioxide that can be produced by the reaction of 95.6 g carbon disulfide with 100. g oxygen?



$$\frac{95.6\text{ g CS}_2}{76.2\text{ g CS}_2} \times \frac{1\text{ mol CS}_2}{1\text{ mol CS}_2} \times \frac{2\text{ mol SO}_2}{1\text{ mol CS}_2} \times \frac{64.1\text{ g SO}_2}{1\text{ mol SO}_2} = 161\text{ g SO}_2$$

$$\frac{100.\text{ g O}_2}{32.0\text{ g O}_2} \times \frac{1\text{ mol O}_2}{1\text{ mol O}_2} \times \frac{2\text{ mol SO}_2}{3\text{ mol O}_2} \times \frac{64.1\text{ g SO}_2}{1\text{ mol SO}_2} = 134\text{ g SO}_2$$

134 g of SO_2 can be produced in this reaction.

Limiting Reactant Answers

Terry C. Jones



Limiting Reactant Answers:

Unveiling the Energy of Verbal Artistry: An Psychological Sojourn through **Limiting Reactant Answers**

In a global inundated with displays and the cacophony of fast communication, the profound power and emotional resonance of verbal artistry usually diminish into obscurity, eclipsed by the regular assault of sound and distractions. Yet, located within the musical pages of **Limiting Reactant Answers**, a charming perform of fictional beauty that pulses with raw emotions, lies an wonderful journey waiting to be embarked upon. Composed by a virtuoso wordsmith, that magical opus manuals visitors on a psychological odyssey, delicately exposing the latent potential and profound influence stuck within the complex web of language. Within the heart-wrenching expanse of this evocative analysis, we can embark upon an introspective exploration of the book is main styles, dissect its fascinating publishing model, and immerse ourselves in the indelible effect it leaves upon the depths of readers souls.

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