





### Stoichiometry Question

Consider the synthesis of water produced in the Model. The reaction container is filled with 10.0 g of  $H_2$  and 1.0 g of  $O_2$ .

- a. Which reactant is the degree of progress in the limiting reactant in this case? Show your work. Also indicate that you are given standard quantities in units such as grams, not moles.

$$10.0 \text{ g } H_2 \left( \frac{1 \text{ mol } H_2}{2.016 \text{ g}} \right) = 4.96 \text{ mol } H_2$$

$$1.0 \text{ g } O_2 \left( \frac{1 \text{ mol } O_2}{32.00 \text{ g}} \right) = 0.031 \text{ mol } O_2$$

	$2 H_2 +$	$O_2 \rightarrow$	$2 H_2O$
Before	10.0 g $H_2$ 	1.0 g $O_2$ 	0.00 g $H_2O$
Change 10.0 g $H_2$ 1.0 g $O_2$ 0.00 g $H_2O$ 0.00 g $H_2O$	-10.0 g $H_2$ -1.0 g $O_2$ 0.00 g $H_2O$ 	+1.0 g $O_2$ +1.0 g $O_2$ -1.0 g $O_2$	+10.0 g $H_2O$ +10.0 g $H_2O$ -10.0 g $H_2O$
After	0.00 g $H_2$ 	0.00 g $O_2$	10.0 g $H_2O$

Change in the limiting reactant

$B + C = A$

- b. What mass of water can be produced? Show your work.

$$0.031 \text{ mol } O_2 \left( \frac{2 \text{ mol } H_2O}{1 \text{ mol } O_2} \right) = 0.062 \text{ mol } H_2O$$

- c. Which reactant is present in excess, and what mass of that reactant remains after the reaction is complete?

$H_2$  is present in excess  
initial mass of  $H_2$  remains

$$10.0 \text{ g } H_2 \left( \frac{2.016 \text{ g}}{1 \text{ mol}} \right) = 4.96 \text{ g } H_2$$

# Limiting Reagent Pogil Answers

**B Lingard**



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