

Stoichiometry and Limiting Reactant Quiz

Version B (PLEASE DO NOT WRITE ON THE QUIZ)

Multiple Choice

Identify the choice that best completes the statement or answers the question.

- ____ 1. When the formula equation $\text{Fe}_3\text{O}_4 + \text{Al} \rightarrow \text{Al}_2\text{O}_3 + \text{Fe}$ is correctly balanced the coefficient of Fe is number
a. 9. c. 7.
b. 2. d. 3.
- ____ 2. In the reaction represented by the equation $\text{Al}_2\text{O}_3 \rightarrow \text{Al} + \text{O}_2$ (unbalanced), what is the mole ratio of aluminum to oxygen?
a. 4:3 c. 10:6
b. 2:3 d. 3:4
- ____ 3. The Haber process for producing ammonia commercially is represented by the equation $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$. To completely convert 9.0 mol hydrogen gas to ammonia gas, how many moles of nitrogen gas are required?
a. 1.0 mol c. 2.0 mol
b. 3.0 mol d. 6.0 mol
- ____ 4. For the reaction represented by the equation $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$, how many grams of water are produced from 6.00 mol of hydrogen?
a. 54.0 g c. 108 g
b. 2.00 g d. 6.00 g
- ____ 5. In most chemical reactions the amount of product obtained is
a. equal to the theoretical yield. c. less than the theoretical yield.
b. more than the theoretical yield. d. more than the percentage yield.
- ____ 6. If the percentage yield is equal to 100%, then
a. there was no limiting reactant.
b. the actual yield is greater than the theoretical yield.
c. the actual yield is less than the theoretical yield.
d. the actual yield is equal to the theoretical yield.
- ____ 7. When two substances react to form products, the reactant which is used up is called the _____.
a. catalytic reagent
b. determining reagent
c. excess reagent
d. limiting reagent
- ____ 8. Which statement is true if 12 mol CO and 12 mol Fe_2O_3 are allowed to react?
 $3\text{CO}(\text{g}) + \text{Fe}_2\text{O}_3(\text{s}) \rightarrow 2\text{Fe}(\text{s}) + 3\text{CO}_2(\text{g})$
a. The limiting reagent is Fe_2O_3 and 36 mol CO_2 will be formed.
b. The limiting reagent is CO and 8.0 mol Fe will be formed.
c. The limiting reagent is Fe_2O_3 and 24 mol Fe will be formed.
d. The limiting reagent is CO and 3.0 mol CO_2 will be formed.

Limiting Reagent Exam Stoichiometry

Kathleen Armour



Limiting Reagent Exam Stoichiometry:

This book delves into Limiting Reagent Exam Stoichiometry. Limiting Reagent Exam Stoichiometry is a crucial topic that must be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Limiting Reagent Exam Stoichiometry, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Limiting Reagent Exam Stoichiometry
- Chapter 2: Essential Elements of Limiting Reagent Exam Stoichiometry
- Chapter 3: Limiting Reagent Exam Stoichiometry in Everyday Life
- Chapter 4: Limiting Reagent Exam Stoichiometry in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, this book will provide an overview of Limiting Reagent Exam Stoichiometry. This chapter will explore what Limiting Reagent Exam Stoichiometry is, why Limiting Reagent Exam Stoichiometry is vital, and how to effectively learn about Limiting Reagent Exam Stoichiometry.

3. In chapter 2, this book will delve into the foundational concepts of Limiting Reagent Exam Stoichiometry. The second chapter will elucidate the essential principles that must be understood to grasp Limiting Reagent Exam Stoichiometry in its entirety.

4. In chapter 3, the author will examine the practical applications of Limiting Reagent Exam Stoichiometry in daily life. The third chapter will showcase real-world examples of how Limiting Reagent Exam Stoichiometry can be effectively utilized in everyday scenarios.

5. In chapter 4, this book will scrutinize the relevance of Limiting Reagent Exam Stoichiometry in specific contexts. The fourth chapter will explore how Limiting Reagent Exam Stoichiometry is applied in specialized fields, such as education, business, and technology.

6. In chapter 5, this book will draw a conclusion about Limiting Reagent Exam Stoichiometry. This chapter will summarize the key points that have been discussed throughout the book.

The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Limiting Reagent Exam Stoichiometry.

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