

Limiting Reactant Practice 2.3 - Answer Key

Objective: Use stoichiometry to identify the limiting and excess reactants.

Directions: Answer the following limiting reactant problems. Show your work!!

1. A reaction between potassium and oxygen gas produces potassium oxide, represented by the following balanced chemical equation:



- a. If 3.00 g of potassium and 5.00 g of oxygen are mixed, how much potassium oxide is produced?

$$3.00 \text{ g K} \times \left(\frac{1 \text{ mol K}}{39.10 \text{ g K}} \right) \left(\frac{1 \text{ mol O}_2}{2 \text{ mol K}} \right) \left(\frac{2 \text{ mol K}_2\text{O}}{1 \text{ mol O}_2} \right) = \boxed{1.21 \text{ g K}_2\text{O}}$$

$$5.00 \text{ g O}_2 \times \left(\frac{1 \text{ mol O}_2}{32.00 \text{ g O}_2} \right) \left(\frac{1 \text{ mol K}_2\text{O}}{1 \text{ mol O}_2} \right) \left(\frac{2 \text{ mol K}_2\text{O}}{1 \text{ mol O}_2} \right) = 5.00 \text{ g K}_2\text{O}$$

- b. What is the limiting reactant?

Potassium (K)

- c. What is the reactant in excess?

Oxygen (O₂)

2. In a reaction, 12.00 g of aqueous sodium hydroxide reacts with aqueous 22.00 g of copper(II) sulfate.



- a. What mass of sodium sulfate is produced?

$$12.00 \text{ g NaOH} \times \left(\frac{1 \text{ mol NaOH}}{40.00 \text{ g NaOH}} \right) \left(\frac{1 \text{ mol Na}_2\text{SO}_4}{2 \text{ mol NaOH}} \right) \left(\frac{142.04 \text{ g Na}_2\text{SO}_4}{1 \text{ mol Na}_2\text{SO}_4} \right) = 17.10 \text{ g}$$

$$22.00 \text{ g CuSO}_4 \times \left(\frac{1 \text{ mol CuSO}_4}{159.61 \text{ g CuSO}_4} \right) \left(\frac{1 \text{ mol Na}_2\text{SO}_4}{1 \text{ mol CuSO}_4} \right) \left(\frac{142.04 \text{ g Na}_2\text{SO}_4}{1 \text{ mol Na}_2\text{SO}_4} \right) = \boxed{19.87 \text{ g}}$$

- b. What is the limiting reactant?

Copper(II) sulfate (CuSO₄)

- c. What is the reactant in excess?

Sodium Hydroxide (NaOH)

Limiting Reactants Study Guide Answers

Christian Drosten



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Table of Contents Limiting Reactants Study Guide Answers

1. Understanding the eBook Limiting Reactants Study Guide Answers
 - The Rise of Digital Reading Limiting Reactants Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reactants Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Limiting Reactants Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reactants Study Guide Answers
 - Personalized Recommendations
 - Limiting Reactants Study Guide Answers User Reviews and Ratings
 - Limiting Reactants Study Guide Answers and Bestseller Lists

5. Accessing Limiting Reactants Study Guide Answers Free and Paid eBooks
 - Limiting Reactants Study Guide Answers Public Domain eBooks
 - Limiting Reactants Study Guide Answers eBook Subscription Services
 - Limiting Reactants Study Guide Answers Budget-Friendly Options
6. Navigating Limiting Reactants Study Guide Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Limiting Reactants Study Guide Answers Compatibility with Devices
 - Limiting Reactants Study Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Limiting Reactants Study Guide Answers
 - Highlighting and Note-Taking Limiting Reactants Study Guide Answers
 - Interactive Elements Limiting Reactants Study Guide Answers
8. Staying Engaged with Limiting Reactants Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Limiting Reactants Study Guide Answers
9. Balancing eBooks and Physical Books Limiting Reactants Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Limiting Reactants Study Guide Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Limiting Reactants Study Guide Answers
 - Setting Reading Goals Limiting Reactants Study Guide Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Limiting Reactants Study Guide Answers
 - Fact-Checking eBook Content of Limiting Reactants Study Guide Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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