

Limiting Reagent Worksheet #1

Key

1. Given the following reaction: (Balance the equation first!)



- a. If you start with 14.8 g of C_3H_8 and 3.44 g of O_2 , determine the limiting reagent.

$$\frac{14.8\text{g C}_3\text{H}_8}{44.09\text{g C}_3\text{H}_8} \times \frac{1\text{mol C}_3\text{H}_8}{1\text{mol C}_3\text{H}_8} \times \frac{4\text{mol H}_2\text{O}}{1\text{mol C}_3\text{H}_8} \times \frac{18.02\text{g H}_2\text{O}}{1\text{mol H}_2\text{O}} = 24.2\text{g H}_2\text{O}$$

$$\frac{3.44\text{g O}_2}{32.00\text{g O}_2} \times \frac{1\text{mol O}_2}{1\text{mol O}_2} \times \frac{4\text{mol H}_2\text{O}}{5\text{mol O}_2} \times \frac{18.02\text{g H}_2\text{O}}{1\text{mol H}_2\text{O}} = 1.55\text{g H}_2\text{O}$$

$\therefore \text{O}_2$ is limiting

- b. Determine the number of grams of carbon dioxide produced.

$$\frac{3.44\text{g O}_2}{32.00\text{g O}_2} \times \frac{1\text{mol O}_2}{1\text{mol O}_2} \times \frac{3\text{mol CO}_2}{5\text{mol O}_2} \times \frac{44.01\text{g CO}_2}{1\text{mol CO}_2} = 2.81\text{g CO}_2$$

- c. Determine the number of grams of H_2O produced.

$$1.55\text{g H}_2\text{O} \quad (\text{from a.})$$

- d. Determine the number of grams of excess reagent left.

$$\frac{3.44\text{g O}_2}{32.00\text{g O}_2} \times \frac{1\text{mol O}_2}{1\text{mol O}_2} \times \frac{1\text{mol C}_3\text{H}_8}{5\text{mol O}_2} \times \frac{44.09\text{g C}_3\text{H}_8}{1\text{mol C}_3\text{H}_8} = 0.948\text{g C}_3\text{H}_8 \text{ used}$$

$$14.8\text{g C}_3\text{H}_8 - 0.948\text{g C}_3\text{H}_8 = 13.9\text{g C}_3\text{H}_8 \text{ remain}$$

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Table of Contents Limiting Reactant Homework Answers On Chemfiesta

1. Understanding the eBook Limiting Reactant Homework Answers On Chemfiesta
 - The Rise of Digital Reading Limiting Reactant Homework Answers On Chemfiesta
 - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reactant Homework Answers On Chemfiesta
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Limiting Reactant Homework Answers On Chemfiesta
 - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reactant Homework Answers On Chemfiesta

- Personalized Recommendations
 - Limiting Reactant Homework Answers On Chemfiesta User Reviews and Ratings
 - Limiting Reactant Homework Answers On Chemfiesta and Bestseller Lists
5. Accessing Limiting Reactant Homework Answers On Chemfiesta Free and Paid eBooks
 - Limiting Reactant Homework Answers On Chemfiesta Public Domain eBooks
 - Limiting Reactant Homework Answers On Chemfiesta eBook Subscription Services
 - Limiting Reactant Homework Answers On Chemfiesta Budget-Friendly Options
 6. Navigating Limiting Reactant Homework Answers On Chemfiesta eBook Formats
 - ePub, PDF, MOBI, and More
 - Limiting Reactant Homework Answers On Chemfiesta Compatibility with Devices
 - Limiting Reactant Homework Answers On Chemfiesta Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Limiting Reactant Homework Answers On Chemfiesta
 - Highlighting and Note-Taking Limiting Reactant Homework Answers On Chemfiesta
 - Interactive Elements Limiting Reactant Homework Answers On Chemfiesta
 8. Staying Engaged with Limiting Reactant Homework Answers On Chemfiesta
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Limiting Reactant Homework Answers On Chemfiesta
 9. Balancing eBooks and Physical Books Limiting Reactant Homework Answers On Chemfiesta
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Limiting Reactant Homework Answers On Chemfiesta
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Limiting Reactant Homework Answers On Chemfiesta
 - Setting Reading Goals Limiting Reactant Homework Answers On Chemfiesta
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Limiting Reactant Homework Answers On Chemfiesta

- Fact-Checking eBook Content of Limiting Reactant Homework Answers On Chemfiesta
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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