

Limiting Reagents and Percentage Yield Worksheet

1. Consider the reaction



- a) 80.0 grams of iodine(V) oxide, I_2O_5 , reacts with 28.0 grams of carbon monoxide, CO . Determine the mass of iodine I_2 , which could be produced?

80 g I_2O_5	1 mol I_2O_5	1 mol I_2		XS
1	333.8 g I_2O_5	1 mol I_2O_5		
28 g CO	1 mol CO	1 mol I_2	253.8 g I_2	
1	28 g CO	5 mol CO	1 mol I_2	= 50.8 g I_2

- b) If, in the above situation, only 0.160 moles, of iodine, I_2 was produced.
i) what mass of iodine was produced?

0.160 mol I_2	253.8 g I_2	= 40.6 g I_2
	1 mol I_2	

- ii) what percentage yield of iodine was produced.

40.6 g I_2	$\times 100 =$	79.9 %
50.8 g I_2		

2. Zinc and sulphur react to form zinc sulphide according to the equation.



If 25.0 g of zinc and 30.0 g of sulphur are mixed,

- a) Which chemical is the limiting reactant?

25.0 g Zn	1 mol Zn	1 mol ZnS	= 382 mol ZnS	LR
1	65.41 g Zn	1 mol Zn		
30.0 g S	1 mol S	1 mol ZnS	= 934 mol ZnS	XS
1	32.1 g S	1 mol S		

- b) How many grams of ZnS will be formed?

382 mol ZnS	97.51 g ZnS	= 37.2 g ZnS
	1 mol ZnS	

- c) How many grams of the excess reactant will remain after the reaction is over?

$$.934 \text{ mol ZnS} - .382 \text{ mol ZnS} = .552 \text{ mol ZnS}$$

.552 mol ZnS	1 mol S	32.1 g S	= 17.7 g S
	1 mol ZnS	1 mol S	

3.

Limiting Reactant And Percent Yield Answers

Roman Wölfel



Limiting Reactant And Percent Yield Answers:

This book delves into Limiting Reactant And Percent Yield Answers. Limiting Reactant And Percent Yield Answers is an essential topic that must be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Limiting Reactant And Percent Yield Answers, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
 - Chapter 1: Introduction to Limiting Reactant And Percent Yield Answers
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 - Chapter 3: Limiting Reactant And Percent Yield Answers in Everyday Life
 - Chapter 4: Limiting Reactant And Percent Yield Answers in Specific Contexts
 - Chapter 5: Conclusion
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