পরীকারী কপুরা, এখানে বোর্র পরীকা, ক্যাডোট কলেলগর শীর্ষান্ধানীর কলেলগমূহের পরীক্ষা এবং মহেল টেন্টের ব্যাত্তর পূর্ণাল উত্তর অধ্যায়ভিত্তিক দেবছা ময়েছে। এগুলো অনুশীলন করনে কৃষি এ অধ্যায়টি লেকে সেকোনো স্থানশীল জনামূদ্

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- ক, মোলার দ্রবণ কীয়
- "কেলাস পানিত্র লকণের দ্রাকারা, তাপমারা বৃশ্বির সাথে সাথে এক পরীয়ে বৃশ্বি না পেয়ে য়াস পায়"— ব্যাব্যা করে।
- চবশ দৃটি মিশ্রিত করলে, থিপ্রদটি কী পরিমাণ অগ্নীয় বা ফারীয় মবে— বিশ্লেষণ করো।

#### ১ নং প্রয়ের উত্তর

1000 mL চৰণে কোন দ্ৰবের 1 mol দ্রবীভূত থাকলে উপ্ত দ্রবণকে
মোলার দ্রবন বলে।

ত্রলাস পানিত্র গবণের প্রাব্যান্তা তাপমাত্রা বৃশ্বির সাথে সাথে এক পর্যান্ত বৃশ্বির না পেরা ছাস পার। এর কারণ, তাপমাত্রা বৃশ্বির সাথে সাবে প্রব্যান্ত বৃশ্বির বিধান পার। এর কারণ, তাপমাত্রা বৃশ্বির সাথে সাবে প্রব্যান বৃশ্বির পোনে থাকার একটি নির্নিন্দ তাপমাত্রার কেলাস পানি কুর গরুণ নির্নান্ত হয়। তথান প্রাব্যান্ত হাস পোনে থাকে। উদাহরণমূল্প Na<sub>2</sub>SO<sub>4</sub>.10H<sub>2</sub>O এর কথা করা যায়। Na<sub>2</sub>SO<sub>4</sub>.10H<sub>2</sub>O এর পেরে পানিতে প্রবীত্ত হওয়ার সময় বিয়োজন প্রক্রিয়াটি ভাপহারী বিক্রিয়া কিন্তু 32.4°C তাপমাত্রার উপরে Na<sub>2</sub>SO<sub>4</sub>.10H<sub>2</sub>O যখন Na<sub>2</sub>SO<sub>4</sub> এ পরিণত হয় তখন বিয়োজন প্রক্রিয়াটি ভাপোৎপাদী হয়। এ কারণে 32.4°C তাপমাত্রার উপরে তাপমাত্রা কৃশ্বিজনিত কারণে প্রাব্যান্ত হস ঘটে।

#### 191

कामता कानि, w = SMV

-(0.5 × 98 × 0.25) g

= 12.25 g

অতএব, দ্বীভূত H₂SO₄ এর ভর ≒ 12.25 g

উদ্দীপকের A-পাত্রে অবস্থিত,

H<sub>2</sub>SO<sub>4</sub> এর আণবিক ভর,

M = 1 × 2 + 32 + 16 × 4

= 2 + 32 + 64

= 98 g/mol

রবণের ঘনমারা,

S = 0.5 M

রবণের আরতন,

V = 250 mL

= 250

1000 = 0.25L

ফরীভূত H<sub>2</sub>SO<sub>4</sub> এর ভর, w = 1

ভারত ন: H<sub>2</sub>SO<sub>4</sub> এর মোল সংখ্যা —  $\left(0.5 \times \frac{250}{1000}\right)$  med — 0.125 mod — 0.125 mod — 0.125 mod — 0.077 mod — 0.077 mod — 0.077 mod সংঘটিত বিভিন্ন: H<sub>2</sub>SO<sub>4</sub> + Na<sub>2</sub>CO<sub>3</sub> — 0.077 mod — 0.077

পাত্র-∧ এর প্রবংগর সাথে পাত্র-৪ এর প্রবংগর বিভিন্নার ফলে ১১৯ moi H₂SO₄ অবশিন্ট থাকায় মিশ্রিত প্রবংগর প্রকৃতি মধে অঞ্জীত

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ভটা ১২ একজন ছাত্র লাবরেটরিতে X-পশ্বতিতে 10 mg বর জেছ হ পরিমাণ এবং Y-পশ্বতিতে 0.5 g পরিমাণ 2% ₩ HCl এর নমূর হ কাজ করে।

- ক, নোড কাকে বলে?
- খ. Cl<sub>2</sub> অপোলার- ব্যাখ্যা করে।।
- উদ্দীপকের এসিভটির মোলার ঘনমারা নির্ণয় করে।
- য়, উদ্দীপকের কোন পশ্বতি পরিবেশবান্ধর? বিশ্লেষণ করে। ২ নং প্রয়োর উত্তর

নিউরিয়াসের চারনিকে যে নির্দিউ ত্রিমাত্রিক স্থানে কোন ইত্রু পাওয়ার সম্ভাব্যতা শূন্য তাকে নোভ বলে।

সমযোজী বন্ধনে আবন্ধ উভয় পরমাণুর তড়িং কণাছকতা ক্র হলে অর্থাৎ তড়িং কণাব্যকতার পার্থকা শূন্য হলে সমযোজী ক বিশুন্ধ ও অপোলার হয়। আর তড়িং ঝণাব্যকতার পার্থকা 0.5 এর হ মর্লে অপোলার হয়। যেহেত CI এর তড়িং ঝণাব্যকতা 3.0 ৫ ট অপুতে তড়িং ঝণাব্যকতার পার্থকা (3 – 3)= 0, তাই CI<sub>2</sub> ক্রিন্ সমযোজী ও অপোলার যৌগ। উরোধ্য মৌলিক অপুর ক্রেন্তে এর্ণ হা থাকে।

ত্রী উদ্দীপকে বর্ণিত এগিচটি হলো 2%  $\left(\frac{W}{V}\right)$  HCI ।

2%  $\left(\frac{W}{V}\right)$  HCI এর অর্থ ফলো-

.. 1000 mL ... ... ...  $\frac{2 \times 1000}{100}$  g

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C. M. Roebuck

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Excel HSC Chemistry C. M. Roebuck. 2003 Excel HSC Chemistry Jim Stamell, 2011 ISBN 9781741252996 AUTHOR Jim Stamell RRP 39 95 PAGES 428 pp SPECIFICATION Softcover perfect bound 280 mm x 210 mm STATUS New edition PUBLICATION DATE April 2008 The EXCEL HSC Chemistry guide is directly linked to the syllabus with every sin gle dot point of the HSC Chemistry syllabus appearing in the margin of the book You can write in the guide so your study is focused and your notes are structured. This guide comes in a brand new format that makes even better use of your study time up to date covera ge of the core topics plus 3 Option topics Industrial Chemistry Shipwr ecks Corrosion and Conservation and Forensic Chemistry this guide is organised just like the HSC syllabus so the students learn to s ection the theoretical part is under routine headings and the students section the practical part is under headings like First hand Second h and Investigations and Problem Solving %this way you will be ab le to see at a glance what the theoretical and practical work is all main headings in each chapter 1 1 2 1 etc are directly from the syllabus word for word %this way you can easily match the E xcel guide to the syllabus an alphabetical list of all the key definitions and concepts you should know from each chapter %an efficient way of learning all the definitions in one go chapter syllabus checklist with every single dot point listed in checklist form for each chapter %a fantastic way of testing that you know all the work hundreds of key concept questions with answers %questions that test you recall of knowledge in each chapter HSC type quest ions for every section in each chapter with clock icons to tell you how much time you will have to answer the questions in the HSC %this way yo u can test yourself on HSC type questions under HSC type time pressure an examiner maximiser feature ticks to show the mark distribut ion and answers to all HSC type questions %all you need to answ er HSC type questions two sample HSC papers with an examiner m aximiser feature plus answers %not one but two up to date sample papers the Excel syllabus summary notes a detachable secti on at the end of the guide where every single dot point of each chapter is summarised for you% a comprehensive and compact summary of the whole course in 32 pages Environmental Chemistry Jorge G. Ibanez, Margarita Hernandez-Esparza, Carmen Doria-Serrano, Arturo Fregoso-Infante, Mono Mohan Singh, 2011-05-06 Environmental issues are growing in importance to the most important political social legal and economic decisions The book presents chemical analyses of our most pressing waste pollution and resource problems for the undergraduate or graduate student The distinctive holistic approach provides a solid ground in theory as well as a laboratory manual detailing introductory and advanced experimental applications The laboratory procedures are presented at microscale conditions for minimum waste and maximum economy This work fulfills an urgent need for an introductory text in environmental chemistry combining theory and practice and is a valuable tool for preparing the next generation of environmental scientists

*Chemical Thermodynamics* W.J. Rankin,2019-11-11 This book develops the theory of chemical thermodynamics from first principles demonstrates its relevance across scientific and engineering disciplines and shows how thermodynamics can be

used as a practical tool for understanding natural phenomena and developing and improving technologies and products Concepts such as internal energy enthalpy entropy and Gibbs energy are explained using ideas and experiences familiar to students and realistic examples are given so the usefulness and pervasiveness of thermodynamics becomes apparent The worked examples illustrate key ideas and demonstrate important types of calculations and the problems at the end of chapters are designed to reinforce important concepts and show the broad range of applications Most can be solved using digitized data from open access databases and a spreadsheet Answers are provided for the numerical problems A particular theme of the book is the calculation of the equilibrium composition of systems both reactive and non reactive and this includes the principles of Gibbs energy minimization The overall approach leads to the intelligent use of thermodynamic software packages but while these are discussed and their use demonstrated they are not the focus of the book the aim being to provide the necessary foundations Another unique aspect is the inclusion of three applications chapters heat and energy aspects of processing the thermodynamics of metal production and recycling and applications of electrochemistry This book is aimed primarily at students of chemistry chemical engineering applied science materials science and metallurgy though it will be also useful for students undertaking courses in geology and environmental science A solutions manual is available for Cambridge Checkpoints HSC Chemistry 2017-19 Maureen Slade, Roger Slade, 2016-06-29 instructors Thermochemistry in Materials Processing Gabriel Plascencia, David Jaramillo, 2017-02-28 This book provides the reader with some thermochemistry notes The intention is to provide a simple easy to understand text which serves as a complimentary material to more complex books It also provide students and those beginning in the field with several application examples used in different areas of materials processing The book presents fully solved problems some quite often found in major metallurgical operations Excel Revise in a Month HSC Chemistry Will Marchment, 2007 **Research & Development** ,1993-09 Metal-Ligand Interactions in Chemistry, Physics and Biology N. Russo, Dennis R. Salahub, 2012-12-06 Proceedings of the NATO Advanced Study Institute held in Cetraro CS Italy from 1 12 September 1998 Resources in Education ,1982

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