

(ix) The standard potential of the cell in the following reaction is _____



($E_{\text{Cd}}^{\circ} = -0.403 \text{ V}$, $E_{\text{Cu}}^{\circ} = 0.334 \text{ V}$)

(a) -0.737 V

(b) 0.737 V

(c) -0.069 V

(d) 0.069 V

(x) The value of $[\text{H}_3\text{O}^+]$ in mol lit^{-1} of 0.001 M acetic acid solution ($K_a = 1.8 \times 10^{-5}$) is _____

(a) 1.34×10^{-1}

(b) 1.34×10^{-2}

(c) 1.34×10^{-3}

(d) 1.34×10^{-4}

Q.2. Answer the following questions:

[8]

(i) Write the product formed when alkyl halide reacts with silver nitrite.

(ii) Write the name of product formed, when acetone is treated with 2, 4-dinitrophenyl hydrazine.

(iii) Write the name of biodegradable polyamide copolymer.

(iv) Identify the molecularity of following elementary reaction:



(v) What is the action of selenium on magnesium metal?

(vi) Write the name of isomerism in the following complexes:



(vii) Write the name of the alloy used in Fischer Tropsch process in the synthesis of gasoline.

(viii) Henry's law constant for $\text{CH}_3\text{Br}_{(\text{g})}$ is $0.159 \text{ mol dm}^{-3} \text{ bar}^{-1}$ at 25°C . What is solubility of $\text{CH}_3\text{Br}_{(\text{g})}$ in water at same temperature and partial pressure of 0.164 bar ?

SECTION – B

Attempt any EIGHT of the following questions:

[16]

Q.3. Explain pseudo-first order reaction with suitable example.

Q.4. Write the consequences of Schottky defect with reasons.

Q.5. What is the action of following on ethyl bromide:

(i) Na in dry ether

(ii) Mg in dry ether

Q.6. Explain formation of peptide linkage in protein with an example.

Q.7. Derive an expression to calculate molar mass of non volatile solute by osmotic pressure measurement.

Q.8. Explain monodentate and ambidentate ligands with example.

Q.9. Explain the trends in the following atomic properties of group 16 elements:

(i) Atomic radii

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Yijin Wang



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Table of Contents Hsc Board Exam Question Papers 2science

1. Understanding the eBook Hsc Board Exam Question Papers 2science
 - The Rise of Digital Reading Hsc Board Exam Question Papers 2science
 - Advantages of eBooks Over Traditional Books
2. Identifying Hsc Board Exam Question Papers 2science
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Hsc Board Exam Question Papers 2science
 - User-Friendly Interface
4. Exploring eBook Recommendations from Hsc Board Exam Question Papers 2science
 - Personalized Recommendations
 - Hsc Board Exam Question Papers 2science User Reviews and Ratings
 - Hsc Board Exam Question Papers 2science and Bestseller Lists

5. Accessing Hsc Board Exam Question Papers 2science Free and Paid eBooks
 - Hsc Board Exam Question Papers 2science Public Domain eBooks
 - Hsc Board Exam Question Papers 2science eBook Subscription Services
 - Hsc Board Exam Question Papers 2science Budget-Friendly Options
6. Navigating Hsc Board Exam Question Papers 2science eBook Formats
 - ePub, PDF, MOBI, and More
 - Hsc Board Exam Question Papers 2science Compatibility with Devices
 - Hsc Board Exam Question Papers 2science Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Hsc Board Exam Question Papers 2science
 - Highlighting and Note-Taking Hsc Board Exam Question Papers 2science
 - Interactive Elements Hsc Board Exam Question Papers 2science
8. Staying Engaged with Hsc Board Exam Question Papers 2science
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Hsc Board Exam Question Papers 2science
9. Balancing eBooks and Physical Books Hsc Board Exam Question Papers 2science
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Hsc Board Exam Question Papers 2science
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Hsc Board Exam Question Papers 2science
 - Setting Reading Goals Hsc Board Exam Question Papers 2science
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Hsc Board Exam Question Papers 2science
 - Fact-Checking eBook Content of Hsc Board Exam Question Papers 2science
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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