

## Assessment 3

**Chapter Test B****Chapter: Chemical Equations and Reactions**

**Part 1** In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- \_\_\_\_\_ 1. The production of a slightly soluble solid compound in a double-displacement reaction is called the formation of a  
a. gas.  
b. precipitate.  
c. molecular reaction.  
d. halogen.
- \_\_\_\_\_ 2. To balance a chemical equation, it is permissible to adjust the  
a. coefficients.  
b. subscripts.  
c. formulae of the products.  
d. number of products.
- \_\_\_\_\_ 3. In a chemical equation, the symbol (aq) indicates that the substance is  
a. melted.  
b. dissolved in water.  
c. an acid.  
d. insoluble.
- \_\_\_\_\_ 4. The tendency for a replacement reaction to occur depends on the  
a. relative tendency of the two elements to lose electrons.  
b. temperature difference.  
c. relative electronegativity.  
d. relative tendency of the two elements to gain electrons.
- \_\_\_\_\_ 5. The coefficients in a chemical equation  
a. indicate the number of moles of each substance that react.  
b. show the number of grams of each substance that react.  
c. are the molar masses of the substances.  
d. show the relative electronegativity of each atom.
- \_\_\_\_\_ 6. If metal X is more reactive than metal Y in the activity series, then  
a. X will replace ions of Y in solution.  
b. Y will replace ions of X in solution.  
c. Y will displace cations only indirectly.  
d. X will react with cold water.

# Holt Chemistry Answer Key Chap 8

**Michael Seilmaier**

A red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the author's name.

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, 1868 **Journal** Royal Society for the Encouragement of Arts, Manufactures and Commerce, 1869 **Journal of the Society of Arts** Royal Society of Arts (Great Britain), 1868 **Principles of Modern Chemistry** David W. Oxtoby, H. Pat Gillis, Laurie J. Butler, 2016-01-01 Long considered the standard for honors and high level mainstream general chemistry courses PRINCIPLES OF MODERN CHEMISTRY continues to set the standard as the most modern rigorous and chemically and mathematically accurate text on the market This authoritative text features an atoms first approach and thoroughly revised chapters on Quantum Mechanics and Molecular Structure Chapter 6 Electrochemistry Chapter 17 and Molecular Spectroscopy and Photochemistry Chapter 20 In addition the text utilizes mathematically accurate and artistic atomic and molecular orbital art and is student friendly without compromising its rigor End of chapter study aids focus on only the most important key objectives equations and concepts making it easier for students to locate chapter content while applications to a wide range of disciplines such as biology chemical engineering biochemistry and medicine deepen students understanding of the relevance of chemistry beyond the classroom Modern Physics Gary N. Felder, Kenny M. Felder, 2022-09-15 An accessible and pedagogically rich Modern Physics textbook with step by step explanations and extensive resources to support active learning *Sif Chemistry Nl Tb* Rex M. Heyworth, 2007 **Introduction to Modern Chemistry** Edward Florian

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aux soins à l'emploi à l'éducation à la liberté de circulation au logement et à la protection sociale

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