

Chapter Test Chapter 10: Matter and Energy

Matter and Energy

Write each answer in the space provided on the paper at the back of this notebook. Show all work for problems that require calculations.

1. ☐ 1. The following are all units of pressure EXCEPT:
 - a. pascals
 - b. newtons per meter
 - c. dynes per centimeter
 - d. atmospheres
2. ☐ 2. Atmospheric pressure is normally:
 - a. the pressure exerted by the entire atmosphere above you
 - b. the atmospheric pressure that acts on a horizontal surface
 - c. the atmospheric pressure that acts on a vertical surface
 - d. the pressure exerted by the air molecules in a volume
3. ☐ 3. Atmospheric pressure varies by about 1 mm Hg for every 10 m of altitude.
 - a. it varies in direct proportion to altitude
 - b. it varies in inverse proportion to altitude
 - c. it varies in direct proportion to the square of altitude
 - d. it varies in inverse proportion to the square of altitude
4. ☐ 4. Which of the following is NOT a unit of pressure?
 - a. newtons per square meter
 - b. dynes per square centimeter
 - c. newtons per square meter
 - d. dynes per square centimeter
5. ☐ 5. Which is a unit of pressure?
 - a. newtons per meter
 - b. newtons per square meter
 - c. newtons per square centimeter
 - d. newtons per square decimeter
6. ☐ 6. The pressure of a gas is 100 kPa. What is the pressure in atm?

$$100 \text{ kPa} \times \frac{1 \text{ atm}}{101.325 \text{ kPa}} = 0.987 \text{ atm}$$
7. ☐ 7. The pressure of a gas is 100 kPa. What is the pressure in mm Hg?

$$100 \text{ kPa} \times \frac{760 \text{ mm Hg}}{101.325 \text{ kPa}} = 750.01 \text{ mm Hg}$$
8. ☐ 8. The pressure of a gas is 100 kPa. What is the pressure in torr?

$$100 \text{ kPa} \times \frac{760 \text{ torr}}{101.325 \text{ kPa}} = 750.01 \text{ torr}$$

Holt Chemistry Test Solutions Answers

Henry Hudson Nicholson



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Holt Chemistry File ,1998 This reference is a must for students who need extra help reteaching or extra practice The guide moves students through the same concepts as the text but at a slower pace More descriptive detail along with visual algorithms provides a more structured approach Each chapter closes with a large bank of practice problems Book jacket

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