

Worksheet #3—Stoichiometry

Directions: Solve the following stoichiometry problems. Assume all gases are at STP. Don't forget to write out your balanced chemical equation. Make sure to show ALL work to receive FULL CREDIT!

1. Calculate the number of moles of oxygen required to react completely with 14.8 mol Al in the following equation: $4\text{Al}_{(s)} + 3\text{O}_{2(g)} \rightarrow 2\text{Al}_2\text{O}_{3(s)}$

$$\frac{14.8 \text{ moles Al}}{4 \text{ mole Al}} \times \frac{3 \text{ moles of O}_2}{1 \text{ mole Al}} = \underline{11.1 \text{ mol O}_2}$$

3. If 20.0g grams of zinc react with excess hydrochloric acid, how many grams of zinc chloride (ZnCl_2) are produced?



$$\frac{20 \text{ g Zn}}{65 \text{ g Zn}} \times \frac{1 \text{ mol Zn}}{1 \text{ mole Zn}} \times \frac{1 \text{ mole ZnCl}_2}{1 \text{ mole Zn}} \times \frac{136 \text{ g ZnCl}_2}{1 \text{ mol ZnCl}_2}$$

$$= \underline{41.8 \text{ g ZnCl}_2}$$

5. How many grams of copper are required to replace 4.00 grams of silver nitrate which are dissolved in water?



$$\frac{4.0 \text{ g AgNO}_3}{170 \text{ g AgNO}_3} \times \frac{1 \text{ mol AgNO}_3}{1 \text{ mole AgNO}_3} \times \frac{2 \text{ mole AgNO}_3}{1 \text{ mole Cu}} \times \frac{63.5 \text{ g Cu}}{1 \text{ mole Cu}}$$

2. Calculate the moles of oxygen produced from the decomposition of 0.61 moles of potassium chlorate (KClO_3).



$$\frac{0.61 \text{ moles KClO}_3}{2 \text{ mole KClO}_3} \times \frac{3 \text{ moles of O}_2}{1 \text{ mole KClO}_3} = \underline{0.915 \text{ mol O}_2}$$

4. How many grams of chlorine gas must be reacted with excess sodium iodide (NaI) if 10.0 grams of sodium chloride (NaCl) are needed



$$\frac{10 \text{ g NaCl}}{58.5 \text{ g NaCl}} \times \frac{1 \text{ mol NaCl}}{1 \text{ mole NaCl}} \times \frac{1 \text{ mole Cl}_2}{2 \text{ mole NaCl}} \times \frac{71 \text{ g Cl}_2}{1 \text{ mole Cl}_2}$$

$$= \underline{6.07 \text{ g Cl}_2}$$

6. How many moles of hydrogen gas are needed to react with 15.1 g chlorine gas to produce hydrogen chloride gas?



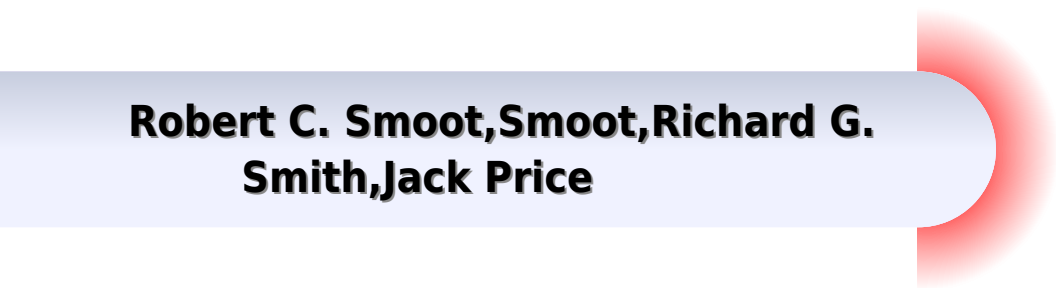
$$\frac{15.1 \text{ g Cl}_2}{71 \text{ g Cl}_2} \times \frac{1 \text{ mol Cl}_2}{1 \text{ mole Cl}_2} \times \frac{1 \text{ mole H}_2}{1 \text{ mole Cl}_2}$$

$$= \underline{0.21 \text{ moles H}_2}$$

Holt Chemistry Problem Solving Work Answers

Stoichiometry

**Robert C. Smoot, Smoot, Richard G.
Smith, Jack Price**



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Brady,John R. Holum,1995-12-29 Offers accurate lucid and interesting explanations of basic concepts and facts of chemistry while helping readers develop skills in analytical thinking and problems solving Modern University Chemistry, Custom Pub Norbert T. Porile,1993 **Chemistry** John W. Moore,Conrad L. Stanitski,Peter C. Jurs,2002 CHEMISTRY THE MOLECULAR SCIENCE is intended to help students develop a broad overview of chemistry and chemical reactions an understanding of the most important concepts and models that chemists and those in chemistry related fields use an appreciation of the many ways chemistry impacts our daily lives the ability to apply the facts concepts and models of chemistry appropriately to new situations in chemistry other sciences and engineering and to other disciplines

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Tzimopoulos,1993 **Chemistry** Jerry A. Bell,American Chemical Society,2005 The American Chemical Society has launched an activities based student centered approach to the general chemistry course a textbook covering all the traditional general chemistry topics but arranged in a molecular context appropriate for biology environmental and engineering students Written by industry chemists and educators Chemistry combines cooperative learning strategies and active learning techniques with a powerful media supplements package to create an effective introductory text Online description **Journal of Chemical Education** ,1926 Includes Report of New England Association of Chemistry Teachers and Proceedings of the Pacific Southwest Association of Chemistry Teachers **Popular Science News** ,1901 **Journal of the American Medical Association** ,1904 **Merrill Chemistry** Robert C. Smoot,Smoot,Richard G. Smith,Jack Price,1998 **Report of the New Jersey State College for the Benefit of Agriculture and the Mechanic Arts, for the Year ...** Rutgers Scientific School,1892 **Resources in education** ,1984-06 Understanding Essential Chemistry Max Diem,2025-05-09 Enables students to understand apply and retain key concepts in general chemistry Understanding

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Table of Contents Holt Chemistry Problem Solving Work Answers Stoichiometry

1. Understanding the eBook Holt Chemistry Problem Solving Work Answers Stoichiometry
 - The Rise of Digital Reading Holt Chemistry Problem Solving Work Answers Stoichiometry
 - Advantages of eBooks Over Traditional Books
2. Identifying Holt Chemistry Problem Solving Work Answers Stoichiometry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Holt Chemistry Problem Solving Work Answers Stoichiometry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Holt Chemistry Problem Solving Work Answers Stoichiometry
 - Personalized Recommendations
 - Holt Chemistry Problem Solving Work Answers Stoichiometry User Reviews and Ratings
 - Holt Chemistry Problem Solving Work Answers Stoichiometry and Bestseller Lists
5. Accessing Holt Chemistry Problem Solving Work Answers Stoichiometry Free and Paid eBooks
 - Holt Chemistry Problem Solving Work Answers Stoichiometry Public Domain eBooks
 - Holt Chemistry Problem Solving Work Answers Stoichiometry eBook Subscription Services

- Holt Chemistry Problem Solving Work Answers Stoichiometry Budget-Friendly Options
- 6. Navigating Holt Chemistry Problem Solving Work Answers Stoichiometry eBook Formats
 - ePub, PDF, MOBI, and More
 - Holt Chemistry Problem Solving Work Answers Stoichiometry Compatibility with Devices
 - Holt Chemistry Problem Solving Work Answers Stoichiometry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Holt Chemistry Problem Solving Work Answers Stoichiometry
 - Highlighting and Note-Taking Holt Chemistry Problem Solving Work Answers Stoichiometry
 - Interactive Elements Holt Chemistry Problem Solving Work Answers Stoichiometry
- 8. Staying Engaged with Holt Chemistry Problem Solving Work Answers Stoichiometry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Holt Chemistry Problem Solving Work Answers Stoichiometry
- 9. Balancing eBooks and Physical Books Holt Chemistry Problem Solving Work Answers Stoichiometry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Holt Chemistry Problem Solving Work Answers Stoichiometry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Holt Chemistry Problem Solving Work Answers Stoichiometry
 - Setting Reading Goals Holt Chemistry Problem Solving Work Answers Stoichiometry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Holt Chemistry Problem Solving Work Answers Stoichiometry
 - Fact-Checking eBook Content of Holt Chemistry Problem Solving Work Answers Stoichiometry
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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