

Acids & Bases Study Guide

6. Water is considered neutral. Describe what this means in terms of $[H^+]$ & $[OH^-]$ concentrations.

8. Complete the following:

$$pH = -\log[H^+]$$

$$K_w = [H^+][OH^-]$$

7. Fill in the blank:

9. What is the $[OH^-]$ of a solution with a pH of 10?

10. What is the $[H^+]$ of a solution with a pH of 4?

11. What is the pH of a solution with a $[H^+]$ of 1×10^{-5} M?

12. What is the pH of a solution with a $[OH^-]$ of 1×10^{-4} M?

13. What is the pH of a solution with a $[H^+]$ of 1×10^{-12} M?

14. What is the pH of a solution with a $[OH^-]$ of 1×10^{-10} M?

15. What is the pH of a solution with a $[H^+]$ of 1×10^{-11} M?

16. What is the pH of a solution with a $[OH^-]$ of 1×10^{-13} M?

Name: _____ Period: _____ Date: _____

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1. Fill in the blanks or complete and contrast the properties of Acids & Bases.

Acids	Bases
<ul style="list-style-type: none"> • Corrosive _____ • Protons _____ • Tastes _____ • Turn litmus paper _____ • High $[H^+]$ _____ 	<ul style="list-style-type: none"> • Corrosive _____ • Protons _____ • Tastes _____ • Turn litmus paper _____ • High $[OH^-]$ _____

2. Fill in the chart for the following acids & bases.

Acid or Base?	Strong or Weak?	Formula	Name
			hydrochloric acid
			sulfuric acid
		HNO_3	
		$NaOH$	
		$Ca(OH)_2$	hydroxide acid

3. Predict the products for the following double replacement reactions involving acids & bases. Identify which component is the Acid, Base, Conjugate Acid (CA), and Conjugate Base (CB) in the formula.



4. Label the pH scale from 0 to 14. Acidic, Basic, Neutral, Strong Acid, Weak Acid, Strong Base, Weak Base.



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