

Name: _____

Chemistry Practice: Balancing Chemical Equations

1. Balance the equations below. Remember this means only using coefficients. All formulas are written correctly.

- a.) $2 \text{NaHCO}_3 \longrightarrow \text{Na}_2\text{CO}_3 + \text{H}_2\text{O} + \text{CO}_2$
- b.) $\text{Zn} + \text{S} \longrightarrow \text{ZnS}$
- c.) $2 \text{H}_2 + \text{O}_2 \longrightarrow 2 \text{H}_2\text{O}$
- d.) $2 \text{FeO} + \text{C} \longrightarrow 2 \text{Fe} + \text{CO}_2$
- e.) $\text{Cl}_2 + 2 \text{KI} \longrightarrow 2 \text{KCl} + \text{I}_2$
- f.) $\text{B}_2\text{O}_3 + 3 \text{Mg} \longrightarrow 3 \text{MgO} + 2 \text{B}$
- g.) $\text{Ca} + 2 \text{H}_2\text{O} \longrightarrow \text{Ca(OH)}_2 + 2 \text{H}_2$
- h.) $4 \text{H}_2\text{O} + \text{Na}_2\text{B}_4\text{O}_7 + \text{H}_2\text{SO}_4 \longrightarrow 4 \text{H}_3\text{BO}_3 + \text{Na}_2\text{SO}_4$
- i.) $\text{CaC}_2 + 2 \text{H}_2\text{O} \longrightarrow \text{Ca(OH)}_2 + \text{C}_2\text{H}_2$
- j.) $\text{NH}_4\text{NO}_3 \longrightarrow \text{N}_2\text{O} + 2 \text{H}_2\text{O}$
- k.) $\text{H}_2\text{S} + \text{O}_2 \longrightarrow \text{SO}_2 + \text{H}_2\text{O}$
- l.) $\text{H}_2\text{SO}_4 + 2 \text{NaCl} \longrightarrow 2 \text{HCl} + \text{Na}_2\text{SO}_4$
- m.) $2 \text{NaCl} + 2 \text{H}_2\text{O} \longrightarrow 2 \text{NaOH} + \text{Cl}_2 + \text{H}_2$
- n.) $3 \text{BaO} + 2 \text{Al} \longrightarrow 3 \text{Ba} + \text{Al}_2\text{O}_3$
- o.) $3 \text{O}_2 \longrightarrow 2 \text{O}_3$
- p.) $\text{SiO}_2 + 2 \text{C} \longrightarrow 2 \text{CO} + \text{Si}$
- q.) $2 \text{PbS} + 3 \text{O}_2 \longrightarrow 2 \text{PbO} + 2 \text{SO}_2$
- r.) $2 \text{HgO} \longrightarrow 2 \text{Hg} + \text{O}_2$
- s.) $\text{C}_{12}\text{H}_{22}\text{O}_{11} + 12 \text{O}_2 \longrightarrow 12 \text{CO}_2 + 11 \text{H}_2\text{O}$
- t.) $\text{C}_3\text{H}_8 + 5 \text{O}_2 \longrightarrow 3 \text{CO}_2 + 4 \text{H}_2\text{O}$
- u.) $2 \text{FeCl}_3 + 3 (\text{NH}_4)_2\text{S} \longrightarrow 6 \text{NH}_4\text{Cl} + \text{Fe}_2\text{S}_3$
- v.) $\text{P}_2\text{O}_5 + 6 \text{NaOH} \longrightarrow 2 \text{Na}_3\text{PO}_4 + 3 \text{H}_2\text{O}$
- w.) $3 \text{BaCl}_2 + \text{Al}_2(\text{SO}_4)_3 \longrightarrow 3 \text{BaSO}_4 + 2 \text{AlCl}_3$
- x.) $4 \text{NH}_3 + 5 \text{O}_2 \longrightarrow 6 \text{H}_2\text{O} + 4 \text{NO}$
- y.) $2 \text{Na}_2\text{O}_2 + 2 \text{H}_2\text{O} \longrightarrow 4 \text{NaOH} + \text{O}_2$
- z.) $\text{Sn(ClO}_3)_4 \longrightarrow \text{SnCl}_4 + 6 \text{O}_2$

Text Problems to be Completed on the Other Side of this Worksheet:

*P. 173 #45,49,53,54,56,57,58,63, 65,66,67,71

*(Although the directions may say to write an "unbalanced" equation, write a balanced equation for each.)

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S Marginson



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Table of Contents Grade 10 Chemistry Exemplar 201

1. Understanding the eBook Grade 10 Chemistry Exemplar 201
 - The Rise of Digital Reading Grade 10 Chemistry Exemplar 201
 - Advantages of eBooks Over Traditional Books
2. Identifying Grade 10 Chemistry Exemplar 201
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Grade 10 Chemistry Exemplar 201
 - User-Friendly Interface
4. Exploring eBook Recommendations from Grade 10 Chemistry Exemplar 201
 - Personalized Recommendations
 - Grade 10 Chemistry Exemplar 201 User Reviews and Ratings
 - Grade 10 Chemistry Exemplar 201 and Bestseller Lists
5. Accessing Grade 10 Chemistry Exemplar 201 Free and Paid eBooks
 - Grade 10 Chemistry Exemplar 201 Public Domain eBooks
 - Grade 10 Chemistry Exemplar 201 eBook Subscription Services
 - Grade 10 Chemistry Exemplar 201 Budget-Friendly Options
6. Navigating Grade 10 Chemistry Exemplar 201 eBook Formats

- ePub, PDF, MOBI, and More
- Grade 10 Chemistry Exemplar 201 Compatibility with Devices
- Grade 10 Chemistry Exemplar 201 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Grade 10 Chemistry Exemplar 201
 - Highlighting and Note-Taking Grade 10 Chemistry Exemplar 201
 - Interactive Elements Grade 10 Chemistry Exemplar 201
- 8. Staying Engaged with Grade 10 Chemistry Exemplar 201
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Grade 10 Chemistry Exemplar 201
- 9. Balancing eBooks and Physical Books Grade 10 Chemistry Exemplar 201
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Grade 10 Chemistry Exemplar 201
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Grade 10 Chemistry Exemplar 201
 - Setting Reading Goals Grade 10 Chemistry Exemplar 201
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Grade 10 Chemistry Exemplar 201
 - Fact-Checking eBook Content of Grade 10 Chemistry Exemplar 201
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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web 1 practice problem balance the redox reaction in basic solution $\text{NO}_2^{-} + \text{Al} + \text{S} + \text{NH}_3 \rightarrow \text{AlO}_2^{-} + \text{S} + \text{N}_2$ practice problem balance the following reduction oxidation reaction in basic solution what is the coefficient in front of OH^{-} $\text{Br}_2 + \text{I}^{-} \rightarrow \text{Br}^{-} + \text{I}_2$ previous topic next topic

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web balance this redox reaction by the half reaction method by assuming an acidic solution $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+}$ balance the redox reaction in exercise 9 by the half reaction method by assuming a basic solution

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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^{-} + \text{Au} + \text{CN}^{-} + 2\text{e}^{-} + 2\text{H}^{+} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2\text{CN}^{-}$ make electrons equal convert to basic solution $4\text{CN}^{-} + 2\text{Au} + 2\text{OH}^{-} + 2\text{e}^{-} \rightarrow 2\text{Au} + 2\text{CN}^{-} + 2\text{H}_2\text{O} + \text{O}_2$ multiplied by a factor of 2 $2\text{H}_2\text{O} + \text{O}_2 \rightarrow 2\text{H}_2\text{O}_2 + 2\text{OH}^{-}$ the final answer

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web aug 29 2023 the first step to balance any redox reaction is to separate the reaction into half reactions the substance being reduced will have electrons as reactants and the oxidized substance will have electrons as products usually all reactions are written as reduction reactions in half reaction tables

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance c e o by adding c e h₂o

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web bonus problem here is a balanced net ionic equation $3\text{Br}_2 + 6\text{OH}^- \rightarrow 5\text{Br}^- + \text{BrO}_3^- + 3\text{H}_2\text{O}$ the above equation takes place in two stages first bromine reacts with oh ions to form bromide ions and bro ions second the bro ions react to form bromide ions and bromate v ions

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1 $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} + \text{Cr}^{3+} + \text{Fe}^{3+}$ solution 1 balanced half reactions $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$ 2 equalize the electrons $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$ multiplied by a factor of 6

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web when balancing equations for redox reactions occurring in acidic solution it is often necessary to add h ions or the h₂o pair to fully balance the equation in this video we ll walk through this process for the reaction between dichromate $\text{Cr}_2\text{O}_7^{2-}$ and chloride cl ions in acidic solution created by jay

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web balance each half reaction in basic solution $4\text{Fe} + 6\text{H}_2\text{O} + 7\text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{NO} + \text{SO}_4^{2-} + 2\text{MnO}_2 + \text{Cr}_3 + \text{NO}_3 + \text{SO}_2 + \text{Mn}_2\text{O}_3$ balance each redox

reaction in acid solution using the half reaction method 8 h₂o₂ cr₂o₇²⁻ 2 9 teo₃ 2 n₂o₄ o₂ cr₃ te 10 reo₄ 11 pbo₂ io i₂ pb₂ 12 as no₃ io₃ re io₃ h₂aso₄

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web practice exercises balance the following equations of redox reactions assign oxidation numbers to all elements in the reaction separate the redox reaction into two half reactions balance the atoms in each half reaction add the two half reactions together and cancel out common terms

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web aug 29 2023 some points to remember when balancing redox reactions the equation is separated into two half equations one for oxidation and one for reduction the equation is balanced by adjusting coefficients and adding h⁺ o²⁻ h⁺ and e⁻ in this order balance the atoms in the equation apart from o and h

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively c 6 h 8 o 6 2h 2no 2 c 6 h 6 o 6 2h 2 o 2no choose 1 answer c 6 h 8 o 6 and no 2 a c 6 h 8 o 6 and no 2 no 2 and c 6 h 8 o 6 b no 2 and c 6 h 8 o 6 h and no 2 c h and