

12.2**FACTORS AFFECTING GAS PRESSURE****SECTION REVIEW****Objectives**

- Explain how the amount of gas and the volume of the container affect gas pressure.
- Infer the effect of temperature changes on the pressure exerted by a contained gas.

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms a chapter has introduced. In each sentence, insert the word or phrase which is most *likely* to make the phrase, clause, or sentence true.

- As long as the temperature of an enclosed gas does not change, _____ 1. _____, doubling the number of gas particles _____ 2. _____ the pressure. When gas particles are removed from a container the pressure decreases. _____ 3. _____ Reducing half the gas particles decreases the pressure by _____ 4. _____.
- When the number of gas particles remains constant, the _____ 5. _____ pressure can be increased by _____ 6. _____ the volume. Reducing the volume of a contained gas by half _____ 7. _____ the pressure.
- When the number of gas particles and the volume of a gas remain constant, the pressure of a gas can be increased by increasing the _____ 8. _____. If the Kelvin temperature doubles, the pressure of the contained gas _____ 9. _____. If the pressure decreases by half when the Kelvin temperature is halved,

Part B True-False

Classify each of these statements as true always true, AT; sometimes true, ST; or never true, NT.

- _____ 1. Halving the number of particles in a given volume of gas decreases the pressure by one-half if the temperature is kept constant.
- _____ 2. Reducing the volume of a container of gas by one-half has the same effect on pressure as doubling the quantity of the gas in the container by one-half, if temperature is kept constant.
- _____ 3. When a sealed container of gas under pressure is opened, gas particles move from the container to the vacuum of air.

Gases And Pressure Section Review Answers

Camilla Rothe



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