

13. Complete the following table for the three experiments in Model 2.

	Experiment C	Experiment D	Experiment E
Independent Variable	# of particles	Temp.	Internal External Pressure
Dependent Variable	Volume	Volume	Volume
Controlled Variable(s)	Temp / I+E pressure	External + Internal Pressure / # of particles	Temp / # of particles

14. Provide a molecular level explanation for the increase in volume among the balloons in Experiment C. (How often and/or how hard are the molecules hitting the sides of the container?)

# of molecules  $\uparrow$ , so number of collisions  $\uparrow$  + Volume along with it

collisions  $\uparrow$  energy stays the same

15. Provide a molecular level explanation for the increase in volume among the balloons in Experiment D.

Temperature  $\uparrow$  increasing the kinetic energy of the particles as well  
which  $\uparrow$  collisions thus  $\uparrow$  pressure

16. Provide a molecular level explanation for the increase in volume among the balloons in Experiment E.

if external forces decrease then it allows internal forces  $\uparrow$   
thus  $\uparrow$  volume

# Gas Variables Pogil

**David Kirk**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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