

NAME: Key DATE: \_\_\_\_\_ PERIOD: \_\_\_\_\_

# GAS STOICHIOMETRY PROBLEMS WORKSHEET 1

1. Given the unbalanced decomposition reaction of baking soda:



42.0 grams of baking soda? What volumes of carbon dioxide and water are produced at STP?

$$42.0 \text{ g NaHCO}_3 \times \frac{1 \text{ mol NaHCO}_3}{84.01 \text{ g NaHCO}_3} \times \frac{1 \text{ mol CO}_2}{2 \text{ mol NaHCO}_3} \times \frac{22.4 \text{ L CO}_2}{1 \text{ mol CO}_2} = \boxed{5.6 \text{ L CO}_2}$$

$$42.0 \text{ g NaHCO}_3 \times \frac{1 \text{ mol NaHCO}_3}{84.01 \text{ g NaHCO}_3} \times \frac{1 \text{ mol H}_2\text{O}}{2 \text{ mol NaHCO}_3} \times \frac{22.4 \text{ L H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = \boxed{5.6 \text{ L H}_2\text{O}}$$

2. The catalytic decomposition of hydrogen peroxide is:



Balance the reaction. How many moles of water and oxygen are produced by the decomposition of 68.0 grams of hydrogen peroxide? How many molecules of water and oxygen are produced? How many grams of each product are formed?

$$68 \text{ g H}_2\text{O}_2 \times \frac{1 \text{ mol H}_2\text{O}_2}{34.01 \text{ g H}_2\text{O}_2} \times \frac{2 \text{ mol H}_2\text{O}}{2 \text{ mol H}_2\text{O}_2} = \boxed{2.00 \text{ mol H}_2\text{O}}$$

$$68 \text{ g H}_2\text{O}_2 \times \frac{1 \text{ mol H}_2\text{O}_2}{34.01 \text{ g H}_2\text{O}_2} \times \frac{1 \text{ mol O}_2}{2 \text{ mol H}_2\text{O}_2} = \boxed{1.00 \text{ mol O}_2}$$

$$68 \text{ g H}_2\text{O}_2 \times \frac{1 \text{ mol H}_2\text{O}_2}{34.01 \text{ g H}_2\text{O}_2} \times \frac{2 \text{ mol H}_2\text{O}}{2 \text{ mol H}_2\text{O}_2} \times \frac{1.80 \times 10^{23} \text{ molecules H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = \boxed{3.60 \times 10^{23} \text{ molecules H}_2\text{O}}$$

$$68 \text{ g H}_2\text{O}_2 \times \frac{1 \text{ mol H}_2\text{O}_2}{34.01 \text{ g H}_2\text{O}_2} \times \frac{1 \text{ mol O}_2}{2 \text{ mol H}_2\text{O}_2} \times \frac{1.80 \times 10^{23} \text{ molecules O}_2}{1 \text{ mol O}_2} = \boxed{1.80 \times 10^{23} \text{ molecules O}_2}$$

$$68 \text{ g H}_2\text{O}_2 \times \frac{1 \text{ mol H}_2\text{O}_2}{34.01 \text{ g H}_2\text{O}_2} \times \frac{2 \text{ mol H}_2\text{O}}{2 \text{ mol H}_2\text{O}_2} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = \boxed{36.04 \text{ g H}_2\text{O}}$$

$$68 \text{ g H}_2\text{O}_2 \times \frac{1 \text{ mol H}_2\text{O}_2}{34.01 \text{ g H}_2\text{O}_2} \times \frac{1 \text{ mol O}_2}{2 \text{ mol H}_2\text{O}_2} \times \frac{32.00 \text{ g O}_2}{1 \text{ mol O}_2} = \boxed{31.96 \text{ g O}_2}$$

3. If 156.0 grams of potassium metal reacts with excess water, potassium hydroxide and hydrogen gas are formed? What volume of hydrogen gas, in liters,

Be sure to balance the reaction first.



$$156 \text{ g K} \times \frac{1 \text{ mol K}}{39.09 \text{ g K}} \times \frac{1 \text{ mol H}_2}{2 \text{ mol K}} \times \frac{22.4 \text{ L H}_2}{1 \text{ mol H}_2} = \boxed{44.70 \text{ L H}_2}$$

$$156 \text{ g K} \times \frac{1 \text{ mol K}}{39.09 \text{ g K}} \times \frac{2 \text{ mol KOH}}{2 \text{ mol K}} \times \frac{56.11 \text{ g KOH}}{1 \text{ mol KOH}} = \boxed{87.39 \text{ g KOH}}$$

4. Determine the number of moles of carbon dioxide gas, water, and sodium chloride formed by the reaction of 42.0 grams of sodium bicarbonate (baking soda) reacting with excess hydrochloric acid. The reaction is:



$$42 \text{ g NaHCO}_3 \times \frac{1 \text{ mol NaHCO}_3}{84.01 \text{ g NaHCO}_3} \times \frac{1 \text{ mol CO}_2}{1 \text{ mol NaHCO}_3} = \boxed{0.5 \text{ mol CO}_2}$$

$$42 \text{ g NaHCO}_3 \times \frac{1 \text{ mol NaHCO}_3}{84.01 \text{ g NaHCO}_3} \times \frac{1 \text{ mol H}_2\text{O}}{1 \text{ mol NaHCO}_3} = \boxed{0.5 \text{ mol H}_2\text{O}}$$

$$42 \text{ g NaHCO}_3 \times \frac{1 \text{ mol NaHCO}_3}{84.01 \text{ g NaHCO}_3} \times \frac{1 \text{ mol NaCl}}{1 \text{ mol NaHCO}_3} = \boxed{0.5 \text{ mol NaCl}}$$

# Gas Stoichiometry Study Guide Answer Key

**Chris Sunley**



## Gas Stoichiometry Study Guide Answer Key:

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## **Table of Contents Gas Stoichiometry Study Guide Answer Key**

1. Understanding the eBook Gas Stoichiometry Study Guide Answer Key
  - The Rise of Digital Reading Gas Stoichiometry Study Guide Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Gas Stoichiometry Study Guide Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Gas Stoichiometry Study Guide Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Stoichiometry Study Guide Answer Key
  - Personalized Recommendations
  - Gas Stoichiometry Study Guide Answer Key User Reviews and Ratings
  - Gas Stoichiometry Study Guide Answer Key and Bestseller Lists
5. Accessing Gas Stoichiometry Study Guide Answer Key Free and Paid eBooks
  - Gas Stoichiometry Study Guide Answer Key Public Domain eBooks
  - Gas Stoichiometry Study Guide Answer Key eBook Subscription Services
  - Gas Stoichiometry Study Guide Answer Key Budget-Friendly Options
6. Navigating Gas Stoichiometry Study Guide Answer Key eBook Formats

- ePub, PDF, MOBI, and More
  - Gas Stoichiometry Study Guide Answer Key Compatibility with Devices
  - Gas Stoichiometry Study Guide Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Gas Stoichiometry Study Guide Answer Key
    - Highlighting and Note-Taking Gas Stoichiometry Study Guide Answer Key
    - Interactive Elements Gas Stoichiometry Study Guide Answer Key
  8. Staying Engaged with Gas Stoichiometry Study Guide Answer Key
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Gas Stoichiometry Study Guide Answer Key
  9. Balancing eBooks and Physical Books Gas Stoichiometry Study Guide Answer Key
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Gas Stoichiometry Study Guide Answer Key
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Gas Stoichiometry Study Guide Answer Key
    - Setting Reading Goals Gas Stoichiometry Study Guide Answer Key
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Gas Stoichiometry Study Guide Answer Key
    - Fact-Checking eBook Content of Gas Stoichiometry Study Guide Answer Key
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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