

Practice Test for Unit 3: Ch. 3, and some of 15 and 16: Kinetic Theory of Matter, States of matter, and energy and thermodynamics, and gas laws.

1. The Kinetic Theory of Matter states that all matter is composed of atoms and molecules that are in a constant state of constant random motion.
2. What is Brownian motion, and why does it happen in terms of kinetic theory?
Small pieces of matter in a fluid – flecks of dust, pieces of pollen, smoke particles, etc. – seem to vibrate randomly because they are being constantly hit from all sides by the molecules of the fluid (water, air, etc.)
3. What is thermal expansion, and why does it happen in terms of kinetic theory?
Thermal expansion is the property of most matter that matter expands as heated and contracts as cooled. All matter is in a constant state of motion, the faster they move, the more space they need to move in and push the other molecules away. Thus as solids and liquids (phases of matter with definite volumes) increase temperature, the molecules move more and take up more space, so the substance increases in size.
4. What is diffusion, and why does it happen in terms of kinetic theory?
The quality of matter that substances in materials where atoms are free to flow (gases or liquids), small particles put into the fluid tend to spread out randomly until they are evenly mixed (go from areas of high concentration to low concentration). This happens because both the fluid and the particles are constantly moving, so will be pushed and dragged around by other molecules until they mix evenly.
5. According to the Kinetic Theory what does temperature measure? (don't say "heat")
The average kinetic energy of the atoms and molecules of a substance
6. How does a thermometer work?
When the molecules of the fluid change temperature, say upwards, they move more, this causes them to take up more space (thermal expansion). The fluid in the thermometer has nowhere else to go, so it rises up the tube. The reverse is true for falling temperatures
7. Why can nothing be cooled to below absolute zero?
At absolute zero all molecular motion stops. Since temperature is the average kinetic energy of the atoms and molecules, if there is no motion, there is no kinetic energy. Something can't move at a slower speed than stopped.
8. How does the freezing point of a substance compare to the melting point of the same substance in the same atmospheric conditions? Why is that?
The freezing point and melting point are the same. They are both changing between a solid and a liquid, so it's the same phase change in reverse. The temperature when the atoms have enough energy such that any added energy will just make them flow past each other (melting) must be the same as the temperature the just flowing liquid would need to such that any removed energy would just lock the atoms in place rather than change their temperature (freezing)

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