

# Key

## Ideal Gas Law Worksheet $PV = nRT$

Use the ideal gas law, " $PerV=nRT$ ", and the universal gas constant  $R = 0.0821 \frac{L \cdot atm}{K \cdot mol}$  to solve the following problems:

If pressure is needed in kPa then convert by multiplying by  $101.3 kPa / 1 atm$  to get  
 $R = 8.31 kPa \cdot L / (K \cdot mole)$

- 1) If I have 4 moles of a gas at a pressure of 5.6 atm and a volume of 12 liters, what is the temperature?

$$PV = nRT$$

$$T = \frac{PV}{nR} = \frac{(5.6 atm)(12 L)}{4 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol}}$$

$$T = 204.63 K$$

- 2) If I have an unknown quantity of gas at a pressure of 1.2 atm, a volume of 31 liters, and a temperature of  $87^{\circ}C$ , how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{1.2 atm \cdot 31 L}{0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 360 K}$$

$$n = 1.2586 mol$$

- 3) If I contain 3 moles of gas in a container with a volume of 60 liters and at a temperature of 400 K, what is the pressure inside the container?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{3 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 400 K}{60 L}$$

$$P = 1.642 atm$$

$$or \quad P = 166.29 kPa$$

- 4) If I have 7.7 moles of gas at a pressure of 0.09 atm and at a temperature of  $56^{\circ}C$ , what is the volume of the container that the gas is in?

$$PV = nRT$$

$$V = \frac{nRT}{P} = \frac{7.7 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 327 K}{0.09 atm}$$

$$V = 2310.93 L$$

- 5) If I have 17 moles of gas at a temperature of  $67^{\circ}C$ , and a volume of 88.89 liters, what is the pressure of the gas?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{17 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 390 K}{88.89 L}$$

$$P = 5.34 atm$$

$$or \quad P = 540.61 kPa$$

- 6) If I have an unknown quantity of gas at a pressure of 0.5 atm, a volume of 25 liters, and a temperature of 300 K, how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{0.5 atm \cdot 25 L}{0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 300 K}$$

$$n = 0.5075 mol$$

# Gas Laws Practice Calculations Answers Unit Lesson 01

**Padhraic Smyth**

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

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## **Table of Contents Gas Laws Practice Calculations Answers Unit Lesson 01**

1. Understanding the eBook Gas Laws Practice Calculations Answers Unit Lesson 01
  - The Rise of Digital Reading Gas Laws Practice Calculations Answers Unit Lesson 01
  - Advantages of eBooks Over Traditional Books
2. Identifying Gas Laws Practice Calculations Answers Unit Lesson 01
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Gas Laws Practice Calculations Answers Unit Lesson 01
  - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Laws Practice Calculations Answers Unit Lesson 01
  - Personalized Recommendations
  - Gas Laws Practice Calculations Answers Unit Lesson 01 User Reviews and Ratings
  - Gas Laws Practice Calculations Answers Unit Lesson 01 and Bestseller Lists
5. Accessing Gas Laws Practice Calculations Answers Unit Lesson 01 Free and Paid eBooks
  - Gas Laws Practice Calculations Answers Unit Lesson 01 Public Domain eBooks
  - Gas Laws Practice Calculations Answers Unit Lesson 01 eBook Subscription Services
  - Gas Laws Practice Calculations Answers Unit Lesson 01 Budget-Friendly Options

6. Navigating Gas Laws Practice Calculations Answers Unit Lesson 01 eBook Formats
  - ePub, PDF, MOBI, and More
  - Gas Laws Practice Calculations Answers Unit Lesson 01 Compatibility with Devices
  - Gas Laws Practice Calculations Answers Unit Lesson 01 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Gas Laws Practice Calculations Answers Unit Lesson 01
  - Highlighting and Note-Taking Gas Laws Practice Calculations Answers Unit Lesson 01
  - Interactive Elements Gas Laws Practice Calculations Answers Unit Lesson 01
8. Staying Engaged with Gas Laws Practice Calculations Answers Unit Lesson 01
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Gas Laws Practice Calculations Answers Unit Lesson 01
9. Balancing eBooks and Physical Books Gas Laws Practice Calculations Answers Unit Lesson 01
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Gas Laws Practice Calculations Answers Unit Lesson 01
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Gas Laws Practice Calculations Answers Unit Lesson 01
  - Setting Reading Goals Gas Laws Practice Calculations Answers Unit Lesson 01
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gas Laws Practice Calculations Answers Unit Lesson 01
  - Fact-Checking eBook Content of Gas Laws Practice Calculations Answers Unit Lesson 01
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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