

- The right side is the product side
- Equations must be balanced!!! They should have the same types and numbers of atoms on each side of the yield.
 - Must follow the Law of Conservation of Matter (Mass): matter (mass) can't be created or destroyed
 - The mass of the reactants must equal the mass of the products. Ex: 50 grams of reactants = 50 grams of product
 - To balance an equation you can change only the coefficients not the subscripts!
- Evidences of a chemical reaction are bubbles, change in color, change in state (solid, liquid, and gas), increase or decrease in temperature; odor or a precipitate forms (solid particles that form when liquids are mixed).
- Mixtures: two or more different substances that are physically combined; retain their own properties
 - Heterogeneous: not uniform throughout, ex: pizza, salads
 - Homogeneous: same throughout, ex: ink
- Factors that affect how quickly a reaction occurs: increased temperature will speed up a reaction, the amount of substance present (concentration) and surface area (how much of the substance is exposed)
- Chemicals and Your Health:
 - Some chemicals can lead to various diseases
 - Radioactive elements are known to cause cancer
 - Asbestos causes cancer
 - Alcohol can affect your liver as well as an unborn child
 - Smoking can cause cancer (ex: lung or mouth cancers)
 - DDT (a pesticide) can cause cancer
 - Mercury can impact brain development
 - Several factors can determine how unhealthy a substance is
 - Concentration = amount of a substance in a certain volume
 - Potency = how "strong" the substance is
 - Dose = amount of a chemical given
 - Exposure (time) = how long you are in contact with the substance
 - Individual susceptible = how a person responds to the substance ("your tolerance")
 - Chemicals and their benefits:
 - Medicines
 - Food preservatives
 - Increase crop production: fertilizers, pesticides (kill insects), herbicides (kill weeds)
 - Sanitation (ex: chlorine used to kill bacteria in water)

➤ The Hydrosphere

➤ Freshwater

- Is polar (has a slight charge creating a positive and negative end)
- Is the universal solvent because it will dissolve so many substances
 - Soluble: ability to dissolve in water
 - Solvent: what does the dissolving
 - Solute: what is being dissolved
 - Supersaturated: no more will dissolve into the water
- States (Phases) of matter: solid → (Latent heat of fusion) liquid → (Heat of Vaporization) gas
- Cohesion: water molecules stick to other water molecules
 - Surface tension: water molecules attracted to other water molecules from all sides and below
- Adhesion: water molecules stick to other surfaces
 - Meniscus: slightly curved water line in the graduated cylinder
- Density: amount of mass in a given volume
 - Formula: Mass/Volume
 - Water's density = 1 g/mL or cm^3

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J Rink

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