

Physical Science Standards:

SPS1: Students will investigate our current understanding of the atom.

a. Examine the structure of the atom in terms of

- proton, electron, and neutron locations.
 - Proton has a charge of _____ and is located in the _____ of the atom.
 - Protons are also the elements _____ number of the identity of the element, so protons are found by looking at the atomic number
 - Neutron has a charge of _____ and is located in the _____ of the atom.
 - Neutrons are found by subtracting the _____ Number from the _____ Mass or the BIG #- SMALL #.
 - Electron has a charge of _____ and is located in the _____ of the atom.
 - Electrons are equal to the number of _____ in a neutral atom.
 - The nucleus is made up of these 2 particles, _____ and _____.
 - The electrons are located in this region of the atom, known as the _____.
 - The electron cloud is divided in to energy levels or energy shells, the first level will hold _____ electrons and the second level will hold _____ electrons.
 - The valence shell or valence level is the _____ energy shell of the atom.
- atomic mass and atomic number.
 - Atomic number is equal to the number of _____ in the atom and is also the atoms identity—it's the small number of the periodic table.
 - Atomic mass is equal to the average mass of all the isotopes of a particular element—it the big number on the periodic table.
- atoms with different numbers of neutrons (isotopes).
 - Define Isotope _____.
 - Ex. Carbon-12 and Carbon-14
 - Carbon 12 has _____ P _____ N _____ E
 - Carbon 14 has _____ P _____ N _____ E
- explain the relationship of the proton number to the element's identity.
 - The _____ number is equal to the number of _____ in the atom which is how the element is identified.
 - Ex. Nitrogen has an atom number of _____, which means that is has _____ protons.

molybdenum	← element name
42	← atomic number number of protons (Z)
Mo	← atomic symbol
95.94	← atomic mass A (this is an average mass)



- Number of Protons = Atomic Number
- Number of Electrons = Number of Protons = Atomic Number
- Number of Neutrons = Mass Number - Atomic Number

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SJ Ball



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Table of Contents Eoct Physics Study Guide

1. Understanding the eBook Eoct Physics Study Guide
 - The Rise of Digital Reading Eoct Physics Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Eoct Physics Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Eoct Physics Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Eoct Physics Study Guide
 - Personalized Recommendations

- Eoct Physics Study Guide User Reviews and Ratings
- Eoct Physics Study Guide and Bestseller Lists
- 5. Accessing Eoct Physics Study Guide Free and Paid eBooks
 - Eoct Physics Study Guide Public Domain eBooks
 - Eoct Physics Study Guide eBook Subscription Services
 - Eoct Physics Study Guide Budget-Friendly Options
- 6. Navigating Eoct Physics Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Eoct Physics Study Guide Compatibility with Devices
 - Eoct Physics Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Eoct Physics Study Guide
 - Highlighting and Note-Taking Eoct Physics Study Guide
 - Interactive Elements Eoct Physics Study Guide
- 8. Staying Engaged with Eoct Physics Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Eoct Physics Study Guide
- 9. Balancing eBooks and Physical Books Eoct Physics Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Eoct Physics Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Eoct Physics Study Guide
 - Setting Reading Goals Eoct Physics Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Eoct Physics Study Guide
 - Fact-Checking eBook Content of Eoct Physics Study Guide

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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