

Limiting Reactant- Theoretical and Percentage Yields Key



If a reaction vessel contains 0.15 mol KO₂ and 0.10 mol H₂O, what is the limiting reactant? How many moles of oxygen can be produced?

$$0.15 \text{ mol KO}_2 \times \frac{3 \text{ mol O}_2}{4 \text{ mol KO}_2} = 0.11 \text{ mol O}_2$$

$$0.10 \text{ mol H}_2\text{O} \times \frac{3 \text{ mol O}_2}{2 \text{ mol H}_2\text{O}} = 0.15 \text{ mol O}_2$$

Limiting reactant: KO₂

Maximum or theoretical yield = 0.11 mol O₂



If a reaction vessel contains 10.0 g of sodium chloride and 12.0 g of sulfuric acid, what is the limiting reactant? What is the theoretical yield of hydrochloric acid? How many grams of excess reactant are left unconsumed?



$$10.0 \text{ g NaCl} \times \frac{1 \text{ mol NaCl}}{58.44 \text{ g NaCl}} \times \frac{2 \text{ mol HCl}}{2 \text{ mol NaCl}} \times \frac{36.46 \text{ g HCl}}{1 \text{ mol HCl}} = 6.24 \text{ g HCl}$$

$$12.0 \text{ g H}_2\text{SO}_4 \times \frac{1 \text{ mol H}_2\text{SO}_4}{98.08 \text{ g H}_2\text{SO}_4} \times \frac{2 \text{ mol HCl}}{1 \text{ mol H}_2\text{SO}_4} \times \frac{36.46 \text{ g HCl}}{1 \text{ mol HCl}} = 8.92 \text{ g HCl}$$

Limiting reactant: NaCl

Maximum or theoretical yield = 6.24 g HCl

$$10.0 \text{ g NaCl} \times \frac{1 \text{ mol NaCl}}{58.44 \text{ g NaCl}} \times \frac{1 \text{ mol H}_2\text{SO}_4}{2 \text{ mol NaCl}} \times \frac{98.08 \text{ g H}_2\text{SO}_4}{1 \text{ mol H}_2\text{SO}_4} = 8.39 \text{ g H}_2\text{SO}_4 \text{ required to consume all of the NaCl (Limiting reactant)}$$

12.0 g H₂SO₄ - 8.39 g H₂SO₄ = 3.61 g of excess H₂SO₄ remains after reaction is complete

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Homework Helpers: Chemistry, Revised Edition Greg Curran, 2011-04-15 Homework Helpers Chemistry is a user friendly review book that will make every student or parent trying to help their child feel like he or she has a private Chemistry tutor Concepts are explained in clear easy to understand language and problems are worked out with step by step methods that are easy to follow Each lesson comes with numerous review questions and answer keynotes that explain each correct answer and why it is correct This book covers all of the topics in a typical one year Chemistry curriculum including A systematic approach to problem solving conversions and the use of units Naming compounds writing formulas and balancing chemical equations Gas laws chemical kinetics acids and bases electrochemistry and more While Homework Helpers Chemistry is an excellent review for any standardized Chemistry test including the SAT II its real value is in providing support and guidance during the year's entire course of study

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