1. What volume of hydrogen gas at STP is produced from the reaction of 50.0g of Mg and 75.0 grams of HCl? How much of the excess reagent is left over (in grams)?

$$Mg(s) + 2 HCl(aq) \rightarrow MgCl_2(s) + H_2(g)$$

Do a standard mass to mass problem starting with each reactant

50.0 gMg	1mol Mg	1mol H ₂	22.4 L	_ =	46.1 L H ₂
	24.31g Mg	1mol Mg	1mol H ₂		

HCl is the limiting reactant!!

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ML Morrison

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