

**Limiting Reagent & Percent Yield Practice Worksheet**

1. When copper (II) chloride reacts with sodium nitrate, copper (II) nitrate and sodium chloride are formed.

a. Write the balanced equation for the reaction given above:



- b. If 15 grams of copper (II) chloride react with 20 grams of sodium nitrate, how much sodium chloride can be formed?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.362 \text{ g CuCl}_2} \times \frac{2 \text{ mol NaCl}}{1 \text{ mol CuCl}_2} \times \frac{58.443 \text{ g NaCl}}{1 \text{ mol NaCl}} = 13 \text{ g NaCl}$$

$$20 \text{ g NaNO}_3 \times \frac{1 \text{ mol NaNO}_3}{84.997 \text{ g NaNO}_3} \times \frac{2 \text{ mol NaCl}}{2 \text{ mol NaNO}_3} \times \frac{58.443 \text{ g NaCl}}{1 \text{ mol NaCl}} = 13.8 \text{ g NaCl}$$

**Only 13 g can be formed**

- c. What is the limiting reagent for the reaction in #2? **copper (II) chloride**

- d. How many grams of copper (II) nitrate is formed?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.452 \text{ g CuCl}_2} \times \frac{1 \text{ mol Cu(NO}_3)_2}{1 \text{ mol CuCl}_2} \times \frac{187.467 \text{ g Cu(NO}_3)_2}{1 \text{ mol Cu(NO}_3)_2} = 20.9 \text{ g Cu(NO}_3)_2$$

- e. How much of the excess reagent is left over in this reaction?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.452 \text{ g CuCl}_2} \times \frac{2 \text{ mol NaNO}_3}{1 \text{ mol CuCl}_2} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 19 \text{ g NaNO}_3$$

$$20 \text{ g NaNO}_3 - 19 \text{ g NaNO}_3 = 1 \text{ g NaNO}_3$$

- f. If 11.3 grams of sodium chloride are formed in the reaction described in problem #2, what is the percent yield of this reaction?

$$\text{Percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100 \text{ percent}$$

$$11.3/13.0 \times 100\% = 86.9\%$$

2. When lead (II) nitrate reacts with sodium iodide, sodium nitrate and lead (II) iodide are formed.

a. Balance the following equation:



- b. If I start with 25.0 grams of lead (II) nitrate and 15.0 grams of sodium iodide, how many grams of sodium nitrate can be formed?

$$25.0 \text{ g Pb(NO}_3)_2 \times \frac{1 \text{ mol Pb(NO}_3)_2}{331.214 \text{ g Pb(NO}_3)_2} \times \frac{2 \text{ mol NaNO}_3}{1 \text{ mol Pb(NO}_3)_2} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 12.8 \text{ g NaNO}_3$$

$$15.0 \text{ g NaI} \times \frac{1 \text{ mol NaI}}{149.89 \text{ g NaI}} \times \frac{2 \text{ mol NaNO}_3}{2 \text{ mol NaI}} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 8.51 \text{ g NaNO}_3$$

**8.51 g NaNO<sub>3</sub> can be made**

# Limiting Reagent And Percent Yield Answer Key

**LM Reder**



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### **Table of Contents Limiting Reagent And Percent Yield Answer Key**

1. Understanding the eBook Limiting Reagent And Percent Yield Answer Key
  - The Rise of Digital Reading Limiting Reagent And Percent Yield Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reagent And Percent Yield Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Limiting Reagent And Percent Yield Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reagent And Percent Yield Answer Key
  - Personalized Recommendations
  - Limiting Reagent And Percent Yield Answer Key User Reviews and Ratings
  - Limiting Reagent And Percent Yield Answer Key and Bestseller Lists

5. Accessing Limiting Reagent And Percent Yield Answer Key Free and Paid eBooks
  - Limiting Reagent And Percent Yield Answer Key Public Domain eBooks
  - Limiting Reagent And Percent Yield Answer Key eBook Subscription Services
  - Limiting Reagent And Percent Yield Answer Key Budget-Friendly Options
6. Navigating Limiting Reagent And Percent Yield Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Limiting Reagent And Percent Yield Answer Key Compatibility with Devices
  - Limiting Reagent And Percent Yield Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Limiting Reagent And Percent Yield Answer Key
  - Highlighting and Note-Taking Limiting Reagent And Percent Yield Answer Key
  - Interactive Elements Limiting Reagent And Percent Yield Answer Key
8. Staying Engaged with Limiting Reagent And Percent Yield Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Limiting Reagent And Percent Yield Answer Key
9. Balancing eBooks and Physical Books Limiting Reagent And Percent Yield Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Limiting Reagent And Percent Yield Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Limiting Reagent And Percent Yield Answer Key
  - Setting Reading Goals Limiting Reagent And Percent Yield Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Limiting Reagent And Percent Yield Answer Key
  - Fact-Checking eBook Content of Limiting Reagent And Percent Yield Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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