Name: Date:

Limiting Reagent Worksheet

Answers

Given the following reaction:

$$C_1H_A + \underline{5} O_1 \rightarrow \underline{3} CO_1 + \underline{4} H_1O_2$$

Balance the reaction.

Suppose 14.8 g of C₃H₈ reacts with 3.44 g of O₂. Determine the limiting reagent.

$$14.8 \text{ g C}_{3}H_{8} \times \frac{1 \text{ mol C}_{3}H_{8}}{44.09 \text{ g C}_{3}H_{8}} \times \frac{4 \text{ mol H}_{2}O}{1 \text{ mol C}_{3}H_{8}} \times \frac{18.02 \text{ g H}_{2}O}{1 \text{ mol H}_{2}O} = 24.2 \text{ g H}_{2}O$$

$$3.44 \text{ g O}_{2} \times \frac{1 \text{ mol O}_{2}}{32 \text{ g O}_{2}} \times \frac{4 \text{ mol H}_{2}O}{5 \text{ mol O}_{2}} \times \frac{18.02 \text{ g H}_{2}O}{1 \text{ mol H}_{2}O} = 1.55 \text{ g H}_{2}O$$

Since O, produces less quantity of H,O than C,H, does, it is the limiting reagent.

(b) Determine the amount of CO, produced.

Given the following reaction:

Balance the reaction.

What is the limiting reagent if 2.2 g of Mg reacts with 4.5 L of O, at STP?

2.2 g Mg *
$$\frac{1 \text{ mol Mg}}{24.3 \text{ g Mg}}$$
 * $\frac{2 \text{ mol MgO}}{2 \text{ mol Mg}}$ = 0.0905 mol MgO
4.5 L O₂ * $\frac{1 \text{ mol O}_2}{22.4 \text{ L O}_2}$ * $\frac{2 \text{ mol MgO}}{1 \text{ mol O}_2}$ * = 0.401 mol MgO

Since Mg produces less MgO than O, Mg is the limiting reagent.

Determine the amount of MgO produced in grams.

Limiting Reactant Work Sheet

Karin Nielsen-Saines



Unveiling the Magic of Words: A Overview of "Limiting Reactant Work Sheet"

In a world defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their power to kindle emotions, provoke contemplation, and ignite transformative change is truly awe-inspiring. Enter the realm of "Limiting Reactant Work Sheet," a mesmerizing literary masterpiece penned with a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve in to the book is central themes, examine its distinctive writing style, and assess its profound impact on the souls of its readers.

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