

CHEMISTRY: LIMITING REAGENT Practice Problems

Please show your work on a separate piece of paper!

Assume standard temperature and pressure, STP conditions for all reactions.

1. At high temperatures, sulfur combines with iron to form the brown-black iron (II) sulfide:



In one experiment, 7.62 g of Fe are allowed to react with 8.67 g of S.

- What is the limiting reagent, and what is the reactant in excess?
 - Calculate the mass of FeS formed.
2. Acrylonitrile, $\text{C}_3\text{H}_3\text{N}$, is the starting material for the production of a kind of synthetic fiber (acrylics) and can be made from propylene, C_3H_6 , by reaction with nitric oxide, NO , as follows:



What mass of $\text{C}_3\text{H}_3\text{N}$ can be made when 21.6 g of C_3H_6 react with 21.6 g of nitric oxide?

3. Calculate the percent yield for the reaction: $\text{P}_4_{(s)} + 6 \text{Cl}_2_{(g)} \rightarrow 4 \text{PCl}_3_{(g)}$
if 75.0 g of phosphorus reacts with excess chlorine gas to produce 111.0 g of phosphorus trichloride.
4. Formic acid, HCHO_2 , burns in oxygen to form carbon dioxide and water as follows:



If a 3.15-g sample of formic acid was burned in 2.0 L of oxygen, what volume of carbon dioxide would be produced?

5. Zinc metal reacts with hydrochloric acid to produce zinc chloride and hydrogen gas.
- Balance the following reaction: $\text{Zn}_{(s)} + \text{HCl}_{(aq)} \rightarrow \text{ZnCl}_2_{(aq)} + \text{H}_2_{(g)}$
 - A 3.50-g sample of zinc metal is allowed to react with 2.50 g of hydrochloric acid.

Complete the following table:

reactants/products	Zn (grams)	HCl (grams)	ZnCl ₂ (grams)	H ₂ (L)
Before reaction				
After reaction	1.26 g			

Limiting Reactant Problems Answer Key

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Table of Contents Limiting Reactant Problems Answer Key

1. Understanding the eBook Limiting Reactant Problems Answer Key
 - The Rise of Digital Reading Limiting Reactant Problems Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reactant Problems Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Limiting Reactant Problems Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reactant Problems Answer Key
 - Personalized Recommendations
 - Limiting Reactant Problems Answer Key User Reviews and Ratings
 - Limiting Reactant Problems Answer Key and Bestseller Lists

5. Accessing Limiting Reactant Problems Answer Key Free and Paid eBooks
 - Limiting Reactant Problems Answer Key Public Domain eBooks
 - Limiting Reactant Problems Answer Key eBook Subscription Services
 - Limiting Reactant Problems Answer Key Budget-Friendly Options
6. Navigating Limiting Reactant Problems Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Limiting Reactant Problems Answer Key Compatibility with Devices
 - Limiting Reactant Problems Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Limiting Reactant Problems Answer Key
 - Highlighting and Note-Taking Limiting Reactant Problems Answer Key
 - Interactive Elements Limiting Reactant Problems Answer Key
8. Staying Engaged with Limiting Reactant Problems Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Limiting Reactant Problems Answer Key
9. Balancing eBooks and Physical Books Limiting Reactant Problems Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Limiting Reactant Problems Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Limiting Reactant Problems Answer Key
 - Setting Reading Goals Limiting Reactant Problems Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Limiting Reactant Problems Answer Key
 - Fact-Checking eBook Content of Limiting Reactant Problems Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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