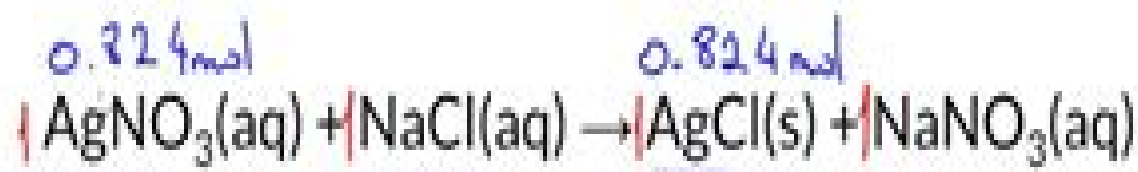


When 140.0 g of  $\text{AgNO}_3$  was added to an aqueous solution of  $\text{NaCl}$ , 86.0 g of  $\text{AgCl}$  was collected as a white precipitate. Which salt was the limiting reactant in this reaction? How many grams of  $\text{NaCl}$  was present in the solution?



1) L.R. - ?

2)  $m(\text{NaCl})$  - ?

$$n(\text{AgNO}_3) = \frac{m}{M} = \frac{140 \text{ g}}{169.9 \text{ g/mol}} = 0.824 \text{ mol}$$

$$n(\text{AgCl}) = \frac{m}{M} = \frac{86.0 \text{ g}}{143.3 \text{ g/mol}} = \underline{0.600 \text{ mol A.Y.}}$$

# Limiting Reactant Practice Problems With Answers

**Julia Burdge**



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In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is actually astonishing. Within the pages of "**Limiting Reactant Practice Problems With Answers**," an enthralling opus penned by a very acclaimed wordsmith, readers embark on an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve into the book's central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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