

ঘের ও কীটপতঙ্গের অংশের নির্ধারিত পটীকার প্রস্তুতকরণ

চলকের ঘনমাত্রা,

$$S = \frac{W}{V} = \frac{20}{36.5} = 1$$

$$M = 0.5479 \text{ M}$$

অতএব, HCl চলকের ঘনমাত্রা

$$= 0.5479 \text{ M}$$

এখানে,

$$W = 20 \text{ g}$$

$$HCl \text{ এর আণবিক ভর,}$$

$$M = 36.5 \text{ g/mol}$$

$$\text{চলকের আয়তন, } V = 1 \text{ L}$$

$$\text{চলকের ঘনমাত্রা } S = 1$$

উদাহরণে উল্লিখিত N_2 -সংশ্লিষ্ট 10 mg এর মধ্যে কয়টি পরিমাণ দ্রব ও V_1 -সংশ্লিষ্ট 0.5 g 2% $\left(\frac{W}{V}\right)$ HCl দ্রবের মধ্যে কাজ করা হয়েছে।

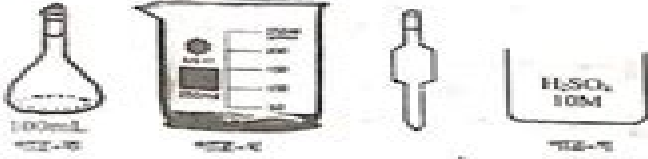
যদিও N_2 -সংশ্লিষ্ট হলো সেমি-মাইক্রো আনালিটিক্যাল পদ্ধতি ও V_1 -সংশ্লিষ্ট হলো ছোট আকারে আনালিটিক্যাল পদ্ধতি। উভয় পদ্ধতির মধ্যে সেমি-মাইক্রো পদ্ধতি পরিবেশ বান্দব। নিম্নে তা বিশ্লেষণ করা হলো-

সেমি-মাইক্রো পদ্ধতি রাসায়নিক গুণগত বিশ্লেষণের একটি পদ্ধতি। এটি বিশ্লেষণের সময় প্রায় 100 মিলিগ্রাম রাসায়নিক উপাদান সন্নিবেশিত ও বিশ্লেষণের বিশ্লেষণে ব্যবহৃত হয়। এ পদ্ধতিতে খুব কম পরিমাণ (0.05 থেকে 0.2g কঠিন বা 2-4ml তরল দ্রব্য) রাসায়নিক পদার্থ নিয়ে পরীক্ষা করা হয়। ফলে পরিবেশে দূষণ উল্লেখযোগ্যভাবে কমে আসে।

এছাড়া সেমি-মাইক্রো পদ্ধতিতে H_2SO_4 এর পরিমাণে কয়েকমিলিগ্রামের ব্যবহার করা হয় ফলে পরিবেশে দূষণ ঘটে না।

$(CH_3NH_2 + H_2O \rightarrow CH_3NH_3^+ + OH^-)$ ফলে, কোনো কোনো ক্ষেত্রে কম কৃত্রিম রাসায়নিক দ্রব্য ব্যবহৃত হয়। ফলে পরিবেশে দূষণের সম্ভাবনা কমে আসে। তাই সেমি-মাইক্রো পদ্ধতি পরিবেশ বান্দব।

চিত্র-১০



১. উদাহরণে দ্রব কী?
২. সেমি-মাইক্রো পদ্ধতি পরিবেশ বান্দব কেন?
৩. চলকের জন্য প্রয়োজ্য হ্যাঙ্গার প্রতীকটি অঙ্কন করে এবং ব্যাখ্যা করে।
৪. উদাহরণের অংশের দ্রব থেকে উপযুক্ত কাচ পাত্র ব্যবহার করে 1M ঘনমাত্রার দ্রব তৈরি করার প্রক্রিয়া ব্যাখ্যা করে। ৪ নং প্রশ্নের উত্তর

১. বিশ্লেষণের নিমিত্ত বীমের উপর রাইডার স্থাপন করলে বীমের প্রতি স্কেলের জন্য যে দূর পাওয়া যায়, তাকে রাইডার দ্রবক বলে।

২. সেমি-মাইক্রো পদ্ধতি রাসায়নিক গুণগত বিশ্লেষণের একটি পদ্ধতি। এটি বিশ্লেষণের সময় প্রায় 100 মিলিগ্রাম রাসায়নিক উপাদান সন্নিবেশিত ও বিশ্লেষণের বিশ্লেষণে ব্যবহৃত হয়। এ পদ্ধতিতে খুব কম পরিমাণ (0.05 থেকে 0.2g কঠিন বা 2-4ml তরল দ্রব্য) রাসায়নিক পদার্থ নিয়ে পরীক্ষা করা হয়। ফলে পরিবেশে দূষণ উল্লেখযোগ্যভাবে কমে আসে।

৩. উদাহরণের দ্রবপটী হলো সালফিউরিক এসিড (H_2SO_4) এর 10 M দ্রব। H_2SO_4 একটি জারক ও ক্ষয়কারী পদার্থ। H_2SO_4 এর 10 M দ্রব দ্রব পাত্র H_2SO_4 দ্রব। পাত্র H_2SO_4 পদার্থকে ক্ষয়কারী হ্যাঙ্গার প্রতীক দিয়ে অঙ্কন করে ব্যাখ্যা করা হলো:

প্রতীক	ক্ষয়কারী (Corrosive)
আয়তন	ক্ষয়কারী রাসায়নিক উপাদান কলতে যে ক্ষয় রাসায়নিক দ্রব্যকে দ্রবের দ্রব সংস্পর্শে তুলে ও অন্যান্য পদার্থের ক্ষয়কারী ক্ষতি সৃষ্টি করে। এতে তুলে পুড়ে যায় এবং অন্যান্য পদার্থ গলে যায়।
উদাহরণ	পাত্র কঠিন এসিডসমূহ (HCl , HNO_3 , H_2SO_4), পাত্র ক্ষয় ($NaOH$, KOH), H_2O_2 , অক্সিজেন লিকার প্রভৃতি দ্রব ইত্যাদি।
সংরক্ষণ পদ্ধতি	- পাত্র এসিডকে পোতলে রাখতে হবে। - পাত্র ক্ষয়কে, ক্ষয়কারী পদার্থের পিছনে রাখতে হবে।

৪. উদাহরণের দ্রবপটী হলো 10M H_2SO_4 দ্রব এবং এর থেকে 1M ঘনমাত্রার দ্রব প্রস্তুত করতে হবে 1ম ও ২ম চলকের ব্যবহার করতে হবে।

1ম চলকের দ্রব হলো 100 ml. অক্সিজেনের দ্রব ও এর দ্রব পিছনে। যেহেতু অক্সিজেনের দ্রবের আয়তন 100 ml. তাই আমাদের 100 ml. 1M ঘনমাত্রার H_2SO_4 দ্রব তৈরি করতে হবে।

$$V_1 S_1 = V_2 S_2$$

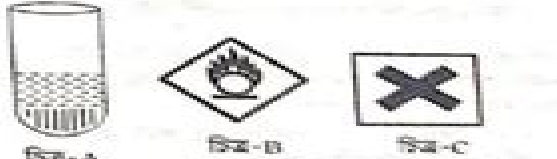
$$\text{বা, } V_1 = \frac{V_2 S_2}{S_1}$$

$$= \left(\frac{100 \times 1}{10} \right) \text{ mL}$$

$$= 10 \text{ mL}$$

অতএব, 10ml H_2SO_4 দ্রবের সঙ্গে পানি যোগ করতে হবে (100-10) mL = 90 mL. আবার, পিছনের দ্রব্যের 10 ml. দ্রব ঘনমাত্রার দ্রব নিয়ে তার সঙ্গে অক্সিজেনের দ্রব 90 ml. পানি যোগ করলে আমাদের কামিষ্ঠ ঘনমাত্রার দ্রব পাওয়া যাবে।

চিত্র-১১



১. অক্সিজেনের দ্রব কী?
২. সোডা দ্রব K_2 এর ঘন কখনো অসীম হয় না-ব্যাখ্যা করে।
৩. A দ্রবের দ্রবের উত্তর করার কৌশল বর্ণনা করে।
৪. B ও C দ্রবের সাথে সম্পর্কিত রাসায়নিক দ্রবের ব্যবহারকালে খাখা কৃতি দ্রবের উপর সম্পর্ক বিশ্লেষণ করে। ৪ নং প্রশ্নের উত্তর

১. অক্সিজেনের দ্রব একটি রাসায়নিক বা কৃত্রিম দ্রব সঞ্চেদ বা দ্রব বা উপাদানসমূহকে জারণ প্রক্রিয়ায় জাল থেকে তুলে করে খাখার সঞ্চেদ করে।

২. সামান্য K_2 এর ঘন কখনো অসীম বা শূন্য হতে পারে না। উদাহরণ ব্যাখ্যা করার জন্য নিম্নের উদাহরণ মাইক্রোপটী বিশ্লেষণ করি-
 $A + B \rightarrow C + D$

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