

Molarity = Moles of solute / Liters of Solution (abbreviation = M)

Molality = Moles of solute / Kg of Solvent (abbreviation = m)

Normality = number of equivalent of solute x Molarity of Solution (abbreviation = N)

Mass Percent = mass of solute / mass of solution

1. How many moles of ethyl alcohol,  $C_2H_5OH$ , are present in 65 mL of a 1.5 M solution?
2. How many liters of a 6.0 M solution of acetic acid  $CH_3COOH$ , contain 0.0030 moles of acetic acid?
3. You want 85 g of KOH. How much of a 3.0 m solution of KOH will provide it?
4. If you dissolve 0.70 moles of HCl in enough water to prepare 250 mL of solution, what is the molarity of the solution you have prepared?
5. A solution is prepared by adding 2.0 L of 6.0 M HCl to 500 mL of a 9.0 M HCl solution. What is the molarity of the new solution? (Remember, the volumes are additive)
6. Convert the following Molarities to Normalities.
  - a. 2.5 M HCl = \_\_\_\_\_ N
  - b. 1.4 M  $H_2SO_4$  = \_\_\_\_\_ N
  - c. 1.0 M NaOH = \_\_\_\_\_ N
  - d. 0.5 M  $Ca(OH)_2$  = \_\_\_\_\_ N
7. A commonly purchased disinfectant is a 3.0% (by mass) solution of hydrogen peroxide ( $H_2O_2$ ) in water. Assuming the density of the solution is 1.0 g/cm<sup>3</sup>, calculate the molarity and molality of  $H_2O_2$ .
8. A solution is made by dissolving 25 g of NaCl in enough water to make 1.0 L of solution. Assume the density of the solution is 1.0 g/cm<sup>3</sup>. Calculate the molarity and molality of the solution.
9. What is the mass percent of a solution that contains 152 g of  $KNO_3$  in 7.86 L of water?
10. A solution has been prepared to be 9.00 % (by mass) glucose. For every 100.0 grams of solution present there are \_\_\_\_\_ grams of glucose.
11. How many grams of KBr are contained in 250. grams of a 6.25 % KBr solution?
12. What is the concentration of each type of ion and total concentration of ions in a 0.375 M Ammonium phosphate solution?
13. How many moles of chloride ions are in 1.50 L of 4.15 M zinc chloride?

**ANSWERS**

1. 0.098 moles of alcohol
2.  $5.0 \times 10^{-4}$  L solution
3.  $5.9 \times 10^{-2}$  g solution
4. 2.8M
5. 6.6 M
6.
  - a. 2.5 N
  - b. 2.8 N
  - c. 1.0 N
  - d. 1 N
7. 0.88M, 0.91m
8. 0.43M, 0.43m
9. 1.90 % by mass of  $KNO_3$
10. 9.00
11. 15.6 g KBr
12.
  - a. 1.13 M of  $(NH_4)^{+1}$  and 0.375 M of  $(PO_4)^{-3}$
  - b. 1.51 M of ions
13. 12.45 moles of  $Cl^{-1}$

# Holt Chemistry Section Concentration And Molarity Answers

**Shasha Hu**



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