# Chapter 2 The Chemistry of Biology

#### Atoms, Bonds, and Molecules

\_\_\_\_\_\_: all tangible materials that occupy space and have mass

Atom: a tiny particle that cannot be subdivided into smaller substances without losing its properties

Properties of atoms:

- Protons (p\*): positively charged, found in the nucleus
- Neutrons (n°): no charge, also found in the nucleus
- Electrons (e'): negatively charged, surround the nucleus

#### **Electron Orbitals and Shells**

Atomic number of protons and neutrons

Orbito)

Pathways in which electrons rotate around the nucleus

## Models of Atomic Structure

in image

#### Major Elements of Life and Their Primary

#### Characteristics 1

Element :

An atom with a characteristic atomic structure and predictable chemical behavior

## **Bonds and Molecules 1**

Molecule

- A chemical substance that results from the combination of two or more atoms
- O<sub>2</sub>, N<sub>2</sub>

Compound

- A molecule composed of two or more different elements
- H<sub>2</sub>O<sub>2</sub> C<sub>6</sub>H<sub>12</sub>O<sub>6</sub>

## Bonds and Molecules 2

Chemical bonds:

Result when two or more atoms share, donate (lose), or accept (gain) electrons

## Covalent Bonds and Polarity 1

Covalent bonds

- Form between atoms that share electrons rather than donating or receiving them
- Single covalent bonds share a pair of electrons (H<sub>2</sub>)
- Double covalent bonds share two pairs of electrons and are more rigid

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Rinehart and Winston Staff,2004 Holt Biology: The environment,2003 Holt Biology Holt Rinehart & Winston,2003-08

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