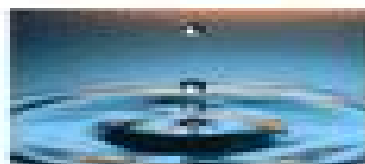


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Intermolecular Forces Lab Worksheet - Answer Key



Directions: Today you will complete two experiments that compare the properties of water to isopropanol (a type of rubbing alcohol). Before conducting each experiment, you will first read about the experiment and then make a prediction about what you think will happen. Afterward you do the experiment, you will write an explanation of what was happening at the atomic level that let us observe these properties.

Introduction: In many ways, water is a miracle liquid. It is essential for all living things on earth and it is often referred to as a universal solvent because many substances dissolve in it. Water displays unusual properties due to the ways the individual water molecules interact with each other.

1. Draw the structural formula (with lone pairs) for the two molecules we will be comparing today, water (H_2O) and isopropanol ($\text{C}_3\text{H}_7\text{O}$). Draw a dipole arrow to show the direction the electrons are being pulled and label the areas of partial positive and negative charges.

H_2O



$\text{C}_3\text{H}_7\text{O}$

Note: the bond between C and H is NOT POLAR. Do not draw dipole arrows or partial charges for those bonds.



2. Which molecule is more polar? water

Experiment #1: Penny

How many drops of water fit on the head of a penny? How many drops of isopropanol fit on the head of a penny?

Prediction:

- a. I think _____ drops of water will fit on the head of the penny because _____

Holt Chemistry Intermolecular Forces Answers

**David W. Oxtoby, H. Pat Gillis, Laurie J.
Butler**



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Holt Chemistry Holt Rinehart & Winston, 2003-01-24 **Holt Chemistry** Ralph Thomas Myers, 2004 **The Handy Chemistry Answer Book** Justin P. Lomont, Ian C. Stewart, 2013-10-01 Don't be mixed up about chemistry Simplify the complex chemical reactions that take place everywhere in our lives with this engaging easy to follow question and answer guide Where would we be without atoms and compounds Gas liquids solids and plasma Acids and bases Bonds and reactions Matter and energy The Handy Chemistry Answer Book covers the building blocks of life and the universe The secret life of atoms how polar bears aren't actually white why oil and water don't mix and much much more are revealed and explained This informative guide covers the basics of chemistry history atomic structures chemical bonds and reactions organic and inorganic chemistry to more advanced material nuclear chemistry biochemistry physical and theoretical chemistry by answering nearly 1 000 common chemistry questions including What causes lightning How does photosynthesis work What are hard and soft Lewis acids and bases What makes a fabric waterproof What are the twelve principles of green chemistry When did alchemists finally abandon trying to make gold What is Le Chatelier's principle What do the different octane ratings mean at the gas pump What is genetic engineering Why is calcium important for strong bones What is the 18 electron rule Why does chocolate turn white as it ages Chemical reactions that rule the world their properties structure composition behavior and history are tackled and explained in plain English in The Handy Chemistry Answer Book With many photos illustrations a few formulas molecular diagrams and other graphics this fun fact filled tome is richly illustrated A history of chemistry timeline appendices on Nobel Prize in Chemistry winners a bibliography further reading section glossary of terms a table of physical constants a table of conversion factors and extensive index add to its usefulness **Principles of Modern Chemistry** David W. Oxtoby, H. Pat Gillis, Laurie J. Butler, 2016-01-01 Long considered the standard for honors and high level mainstream general chemistry courses PRINCIPLES OF MODERN CHEMISTRY continues to set the standard as the most modern rigorous and chemically and mathematically accurate text on the market This authoritative text features an atoms first approach and thoroughly revised chapters on Quantum Mechanics and Molecular Structure Chapter 6 Electrochemistry Chapter 17 and Molecular Spectroscopy and Photochemistry Chapter 20 In addition the text utilizes mathematically accurate and artistic atomic and molecular orbital art and is student friendly without compromising its rigor End of chapter study aids focus on only the most important key objectives equations and concepts making it easier for students to locate chapter content while applications to a wide range of disciplines such as biology chemical engineering biochemistry and medicine deepen students understanding of the relevance of chemistry beyond the classroom

Intermolecular Forces Pierre L. Huyskens, Werner A.P. Luck, Therese Zeegers-Huyskens, 2012-12-06 The study of intermolecular forces began over one hundred years ago in 1873 with the famous thesis of van der Waals In recent decades knowledge of this field has expanded due to intensive research into both its theoretical and the experimental aspects This is

particularly true for the type of very strong cohesive force stressed in 1920 by Latimer and Rodebush the hydrogen bond a phenomenon already outlined in 1912 by Moore and Winemill Hydrogen bonds exert a profound influence on most of the physical and chemical properties of the materials in which they are formed Not only do they govern viscosity and electrical conductivity they also intervene in the chemical reaction path which determines the kinetics of chemical processes The properties of chemical substances depend to a large extent on intermolecular forces In spite of this fundamental fact too little attention is given to these properties both in research and in university teaching For instance in the field of pharmaceutical research about 13000 compounds need to be studied in order to find a single new product that can be successfully marketed The recognition of the need to optimize industrial research efficiency has led to a growing interest in promoting the study of inter molecular forces Rising salary costs in industry have encouraged an interest in theoretical ideas which will lead to tailor made materials

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Modern Chemistry Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001

Fundamentals of Modern Bioprocessing Sarfaraz K. Niazi, Justin L. Brown, 2017-07-27 Biological drug and vaccine manufacturing has quickly become one of the highest value fields of bioprocess engineering and many bioprocess engineers are now finding job opportunities that have traditionally gone to chemical engineers Fundamentals of Modern Bioprocessing addresses this growing demand

Written by experts well established in the field this book connects the principles and applications of bioprocessing engineering to healthcare product manufacturing and expands on areas of opportunity for qualified bioprocess engineers and students The book is divided into two sections the first half centers on the engineering fundamentals of bioprocessing while the second half serves as a handbook offering advice and practical applications Focused on the fundamental principles at the core of this discipline this work outlines every facet of design component selection and regulatory concerns It discusses the purpose of bioprocessing to produce products suitable for human use describes the manufacturing technologies related to bioprocessing and explores the rapid expansion of bioprocess engineering applications relevant to health care product manufacturing It also considers the future of bioprocessing the use of disposable components which is the fastest growing area in the field of bioprocessing to replace traditional stainless steel In addition this text Discusses the many types of genetically modified organisms Outlines laboratory techniques Includes the most recent developments Serves as a reference and contains an extensive bibliography Emphasizes biological manufacturing using recombinant processing which begins with creating a genetically modified organism using recombinant techniques Fundamentals of Modern Bioprocessing outlines both the principles and applications of bioprocessing engineering related to healthcare product manufacturing It lays out the basic concepts definitions methods and applications of bioprocessing A single volume comprehensive reference developed to meet the needs of students with a bioprocessing background it can also be used as a source for professionals in the field

Fundamentals of Chemistry: A Modern Introduction (1966) Frank Brescia, 2012-12-02 Fundamentals of Chemistry A Modern Introduction focuses on the formulas processes and methodologies used in the study of chemistry The book first looks at general and historical remarks definitions of chemical terms and the classification of matter and states of aggregation The text then discusses gases Ideal gases pressure of a gas confined by a liquid Avogadro's Law and Graham's Law are described The book also discusses aggregated states of matter atoms and molecules chemical equations and arithmetic thermochemistry and chemical periodicity The text also highlights the electronic structures of atoms Quantization of electricity spectra of elements quantization of the energy of an electron associated with nucleus the Rutherford Bohr nuclear theory hydrogen atom and representation of the shapes of atomic orbitals are explained The text also highlights the types of chemical bonds hydrocarbons and their derivatives intermolecular forces solutions and chemical equilibrium The book focuses as well on ionic solutions galvanic cells and acids and bases It also discusses the structure and basicity of hydrides and oxides The reactivity of hydrides charge of dispersal and basicity effect of anionic charge inductive effect and basicity and preparation of acids are described The book is a good source of information for readers wanting to study chemistry **Introduction to Modern Chemistry** Edward Florian Neuzil, 1968 **Modern Chemistry**, 1990

Fundamentals of Chemistry: A Modern Introduction Frank Brescia, 2012-12-02 Fundamentals of Chemistry Third Edition introduces the reader to the fundamentals of chemistry including the properties of gases atomic and molecular

weights and the first and second laws of thermodynamics Chemical equations and chemical arithmetic are also discussed along with the structure of atoms chemical periodicity types of chemical bonds and condensed states of matter This book is comprised of 26 chapters and begins with a historical overview of chemistry and some terms which are part of the language of chemists Separation and purification are covered in the first chapter while the following chapters focus on atomic and molecular weights stoichiometry the structure of atoms and types of chemical bonds The molecular orbital MO theory of bonding galvanic cells and chemical thermodynamics are considered next Separate chapters are devoted to MO theory of covalent and metallic bonding orbital hybridization intermolecular forces acids and bases ionic equilibrium calculations and polymers and biochemicals This monograph is intended for chemistry students [MCAT General Chemistry Review](#)

Princeton Review (Firm), 2010 The MCAT is a test of more than just the facts about basic physical and biological sciences it is an in depth rigorous examination of your knowledge of scientific concepts and principles as well as your critical thinking and writing skills With the Princeton Review's subject specific MCAT series you can focus your review on the MCAT topics that are most challenging to you Each book in the series contains the most in depth coverage of subjects tested on the MCAT Each chapter in MCAT General Chemistry includes Full color illustrations charts and diagrams Examples of general chemistry questions and their solutions worked out step by step Chapter Review Quizzes and answers A real MCAT style practice passage with questions and answers Bulleted chapter summaries for quick review MCAT General Chemistry Review also includes A complete glossary of general chemistry terms A general chemistry formula sheet **College Chemistry**

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