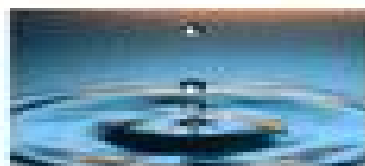


Name: _____

Date: _____

Class: _____

Intermolecular Forces Lab Worksheet - Answer Key



Directions: Today you will complete two experiments that compare the properties of water to isopropanol (a type of rubbing alcohol). Before conducting each experiment, you will first read about the experiment and then make a prediction about what you think will happen. Afterward you do the experiment, you will write an explanation of what was happening at the atomic level that let us observe these properties.

Introduction: In many ways, water is a miracle liquid. It is essential for all living things on earth and it is often referred to as a universal solvent because many substances dissolve in it. Water displays unusual properties due to the ways the individual water molecules interact with each other.

1. Draw the structural formula (with lone pairs) for the two molecules we will be comparing today, water (H_2O) and isopropanol ($\text{C}_3\text{H}_8\text{O}$). Draw a dipole arrow to show the direction the electrons are being pulled and label the areas of partial positive and negative charges.

H_2O



or



$\text{C}_3\text{H}_8\text{O}$

Note: the bond between C and H is NOT POLAR. Do not draw dipole arrows or partial charges for those bonds.



2. Which molecule is more polar? water

Experiment #1: Penny

How many drops of water fit on the head of a penny? How many drops of isopropanol fit on the head of a penny?

Prediction:

- a. I think _____ drops of water will fit on the head of the penny because _____

Holt Chemistry Concept Review Intermolecular Forces

Neil D. Jespersen, Pamela Kerrigan



Holt Chemistry Concept Review Intermolecular Forces:

This book delves into Holt Chemistry Concept Review Intermolecular Forces. Holt Chemistry Concept Review Intermolecular Forces is a vital topic that needs to be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Holt Chemistry Concept Review Intermolecular Forces, encompassing both the fundamentals and more intricate discussions.

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