

Hewlett-Packard Financial Performance Assessment

Exercise 10.17 has three similar following problems. Use all appropriate and proper scientific units. Use scientific notation for answers (e.g., 1.23×10^4).

1. Find the standard enthalpy change for the reaction $\text{C}_{10}\text{H}_{18} + 15\text{O}_2 \rightarrow 10\text{CO}_2 + 9\text{H}_2\text{O}$.
 Given that $\Delta H_{\text{f}}^\circ(\text{C}_{10}\text{H}_{18}) = 125 \text{ kJ/mol}$ and $\Delta H_{\text{f}}^\circ(\text{C}_2\text{H}_6) = 85 \text{ kJ/mol}$
 and $\Delta H_{\text{f}}^\circ(\text{C}_2\text{H}_4) = 52 \text{ kJ/mol}$ and $\Delta H_{\text{f}}^\circ(\text{C}_2\text{H}_2) = 227 \text{ kJ/mol}$ and $\Delta H_{\text{f}}^\circ(\text{H}_2) = 0 \text{ kJ/mol}$

a. $\Delta H_{\text{rxn}}^\circ = \Delta H_{\text{f}}^\circ(\text{C}_{10}\text{H}_{18}) + 15\Delta H_{\text{f}}^\circ(\text{O}_2) - 10\Delta H_{\text{f}}^\circ(\text{CO}_2) - 9\Delta H_{\text{f}}^\circ(\text{H}_2\text{O})$ $\Delta H_{\text{rxn}}^\circ = 125 \text{ kJ/mol}$
 $\Delta H_{\text{rxn}}^\circ = 125 \text{ kJ/mol} + 15(0 \text{ kJ/mol}) - 10(-393 \text{ kJ/mol}) - 9(-285 \text{ kJ/mol})$ $\Delta H_{\text{rxn}}^\circ = 125 \text{ kJ/mol}$

b. $\Delta H_{\text{rxn}}^\circ = 10\Delta H_{\text{f}}^\circ(\text{CO}_2) + 9\Delta H_{\text{f}}^\circ(\text{H}_2\text{O}) - \Delta H_{\text{f}}^\circ(\text{C}_{10}\text{H}_{18}) - 15\Delta H_{\text{f}}^\circ(\text{O}_2)$ $\Delta H_{\text{rxn}}^\circ = 125 \text{ kJ/mol}$

2. Find standard enthalpy change for the formation of acetylene from ethylene and ethane.
 $\text{C}_2\text{H}_4 + 2\text{C}_2\text{H}_6 \rightarrow 2\text{C}_2\text{H}_2 + 2\text{H}_2$ $\Delta H_{\text{rxn}}^\circ = 209 \text{ kJ/mol}$
 $\Delta H_{\text{f}}^\circ(\text{C}_2\text{H}_4) = 52 \text{ kJ/mol}$ and $\Delta H_{\text{f}}^\circ(\text{C}_2\text{H}_6) = 85 \text{ kJ/mol}$ and $\Delta H_{\text{f}}^\circ(\text{C}_2\text{H}_2) = 227 \text{ kJ/mol}$
 Standard enthalpy of formation: $\Delta H_{\text{f}}^\circ(\text{H}_2) = 0 \text{ kJ/mol}$ and $\Delta H_{\text{f}}^\circ(\text{H}_2\text{O}) = -285 \text{ kJ/mol}$

a. $\Delta H_{\text{rxn}}^\circ = 2\Delta H_{\text{f}}^\circ(\text{C}_2\text{H}_2) + 2\Delta H_{\text{f}}^\circ(\text{H}_2) - \Delta H_{\text{f}}^\circ(\text{C}_2\text{H}_4) - 2\Delta H_{\text{f}}^\circ(\text{C}_2\text{H}_6)$ $\Delta H_{\text{rxn}}^\circ = 209 \text{ kJ/mol}$
 $\Delta H_{\text{rxn}}^\circ = 2(227 \text{ kJ/mol}) + 2(0 \text{ kJ/mol}) - 52 \text{ kJ/mol} - 2(85 \text{ kJ/mol})$ $\Delta H_{\text{rxn}}^\circ = 209 \text{ kJ/mol}$

b. $\Delta H_{\text{rxn}}^\circ = 2\Delta H_{\text{f}}^\circ(\text{C}_2\text{H}_2) + 2\Delta H_{\text{f}}^\circ(\text{H}_2) - \Delta H_{\text{f}}^\circ(\text{C}_2\text{H}_4) - 2\Delta H_{\text{f}}^\circ(\text{C}_2\text{H}_6)$ $\Delta H_{\text{rxn}}^\circ = 209 \text{ kJ/mol}$
 $\Delta H_{\text{rxn}}^\circ = 2(227 \text{ kJ/mol}) + 2(0 \text{ kJ/mol}) - 52 \text{ kJ/mol} - 2(85 \text{ kJ/mol})$ $\Delta H_{\text{rxn}}^\circ = 209 \text{ kJ/mol}$

c. $\Delta H_{\text{rxn}}^\circ = 2\Delta H_{\text{f}}^\circ(\text{C}_2\text{H}_2) + 2\Delta H_{\text{f}}^\circ(\text{H}_2) - \Delta H_{\text{f}}^\circ(\text{C}_2\text{H}_4) - 2\Delta H_{\text{f}}^\circ(\text{C}_2\text{H}_6)$ $\Delta H_{\text{rxn}}^\circ = 209 \text{ kJ/mol}$

3. Find Gibbs free energy of reaction for the combustion of carbon monoxide and hydrogen, which would give the equation: $\text{CO} + \text{H}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ and $\Delta H_{\text{rxn}}^\circ = 75 \text{ kJ/mol}$

Standard enthalpy change for the reaction from the following standard enthalpies:

$\Delta H_{\text{f}}^\circ(\text{CO}) = 110 \text{ kJ/mol}$ $\Delta H_{\text{f}}^\circ(\text{CO}_2) = 393 \text{ kJ/mol}$ $\Delta H_{\text{f}}^\circ(\text{H}_2\text{O}) = -285 \text{ kJ/mol}$
 $\Delta H_{\text{f}}^\circ(\text{H}_2) = 0 \text{ kJ/mol}$ $\Delta H_{\text{f}}^\circ(\text{C}_2\text{H}_6) = 85 \text{ kJ/mol}$ $\Delta H_{\text{f}}^\circ(\text{C}_2\text{H}_4) = 52 \text{ kJ/mol}$

a. $\Delta H_{\text{rxn}}^\circ = \Delta H_{\text{f}}^\circ(\text{CO}_2) + \Delta H_{\text{f}}^\circ(\text{H}_2\text{O}) - \Delta H_{\text{f}}^\circ(\text{CO}) - \Delta H_{\text{f}}^\circ(\text{H}_2)$ $\Delta H_{\text{rxn}}^\circ = 75 \text{ kJ/mol}$
 $\Delta H_{\text{rxn}}^\circ = 393 \text{ kJ/mol} + (-285 \text{ kJ/mol}) - 110 \text{ kJ/mol} - 0 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 75 \text{ kJ/mol}$
 $\Delta H_{\text{rxn}}^\circ = 393 \text{ kJ/mol} - 285 \text{ kJ/mol} - 110 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 75 \text{ kJ/mol}$

b. $\Delta H_{\text{rxn}}^\circ = \Delta H_{\text{f}}^\circ(\text{CO}_2) + \Delta H_{\text{f}}^\circ(\text{H}_2\text{O}) - \Delta H_{\text{f}}^\circ(\text{CO}) - \Delta H_{\text{f}}^\circ(\text{H}_2)$ $\Delta H_{\text{rxn}}^\circ = 75 \text{ kJ/mol}$
 $\Delta H_{\text{rxn}}^\circ = 393 \text{ kJ/mol} + (-285 \text{ kJ/mol}) - 110 \text{ kJ/mol} - 0 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 75 \text{ kJ/mol}$
 $\Delta H_{\text{rxn}}^\circ = 393 \text{ kJ/mol} - 285 \text{ kJ/mol} - 110 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 75 \text{ kJ/mol}$

c. $\Delta H_{\text{rxn}}^\circ = \Delta H_{\text{f}}^\circ(\text{CO}_2) + \Delta H_{\text{f}}^\circ(\text{H}_2\text{O}) - \Delta H_{\text{f}}^\circ(\text{CO}) - \Delta H_{\text{f}}^\circ(\text{H}_2)$ $\Delta H_{\text{rxn}}^\circ = 75 \text{ kJ/mol}$

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