

Gas Variables

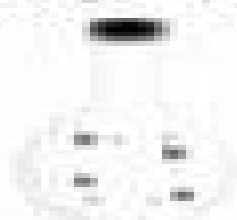
How are the variables that describe a gas related?

Why?

Imagine buying a balloon bouquet at a party store. How will the helium gas in the bouquet behave if you carry it outside on a hot summer day? How will it behave if you carry it outside during a snowstorm? What happens if the balloons are made of latex, which can stretch? What happens if the balloons are made of Mylar[®], which cannot stretch? What if you add just a small amount of gas to each balloon? What if you add a lot of gas? In this activity, you will explore four variables that quantify gases—pressure (P), volume (V), temperature (T), and moles (n) of gas. These four variables can be related mathematically so that predictions about gas behavior can be made.

Model 1 – Gases in a Nonflexible Container

Experiment A (Adding more gas)



A1

Volume = 2 unit

External pressure = 1 atm

Internal pressure = 1 atm

Temperature = 300 K



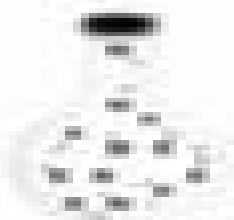
A2

Volume = 2 unit

External pressure = 1 atm

Internal pressure = 2 atm

Temperature = 300 K



A3

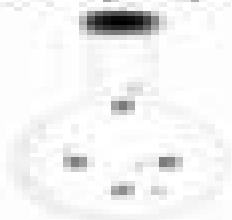
Volume = 2 unit

External pressure = 1 atm

Internal pressure = 3 atm

Temperature = 300 K

Experiment B (Heating the gas)



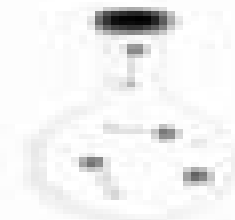
B1

Volume = 2 unit

External pressure = 1 atm

Internal pressure = 1 atm

Temperature = 300 K



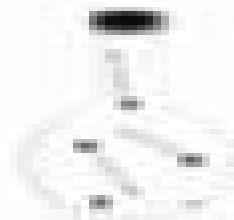
B2

Volume = 2 unit

External pressure = 1 atm

Internal pressure = 2 atm

Temperature = 400 K



B3

Volume = 2 unit

External pressure = 1 atm

Internal pressure = 3 atm

Temperature = 600 K

*Note: Volume in the model is not set in units, rather than liters, because 0 molecules of gas at the conditions given would take every third space (1×10^{-21} m). These particles themselves are much larger compared to the space between them than in a liquid or solid.

Gas Variables Packet

PT Brinkman

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the PT Brinkman text.

Gas Variables Packet:

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Table of Contents Gas Variables Packet

1. Understanding the eBook Gas Variables Packet
 - The Rise of Digital Reading Gas Variables Packet
 - Advantages of eBooks Over Traditional Books
2. Identifying Gas Variables Packet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gas Variables Packet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Variables Packet
 - Personalized Recommendations
 - Gas Variables Packet User Reviews and Ratings
 - Gas Variables Packet and Bestseller Lists
5. Accessing Gas Variables Packet Free and Paid eBooks
 - Gas Variables Packet Public Domain eBooks
 - Gas Variables Packet eBook Subscription Services
 - Gas Variables Packet Budget-Friendly Options

6. Navigating Gas Variables Packet eBook Formats
 - ePub, PDF, MOBI, and More
 - Gas Variables Packet Compatibility with Devices
 - Gas Variables Packet Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gas Variables Packet
 - Highlighting and Note-Taking Gas Variables Packet
 - Interactive Elements Gas Variables Packet
8. Staying Engaged with Gas Variables Packet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gas Variables Packet
9. Balancing eBooks and Physical Books Gas Variables Packet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gas Variables Packet
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gas Variables Packet
 - Setting Reading Goals Gas Variables Packet
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gas Variables Packet
 - Fact-Checking eBook Content of Gas Variables Packet
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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