

NAME: Key DATE: \_\_\_\_\_ PERIOD: \_\_\_\_\_

# GAS STOICHIOMETRY PROBLEMS WORKSHEET 1

1. Given the unbalanced decomposition reaction of baking soda:



42.0 grams of baking soda? What volumes of carbon dioxide and water are produced at STP?

$$42.0 \text{ g NaHCO}_3 \times \frac{1 \text{ mol NaHCO}_3}{84.01 \text{ g NaHCO}_3} \times \frac{1 \text{ mol CO}_2}{2 \text{ mol NaHCO}_3} \times \frac{22.4 \text{ L CO}_2}{1 \text{ mol CO}_2} = \boxed{5.6 \text{ L CO}_2}$$

$$42.0 \text{ g NaHCO}_3 \times \frac{1 \text{ mol NaHCO}_3}{84.01 \text{ g NaHCO}_3} \times \frac{1 \text{ mol H}_2\text{O}}{2 \text{ mol NaHCO}_3} \times \frac{22.4 \text{ L H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = \boxed{5.6 \text{ L H}_2\text{O}}$$

2. The catalytic decomposition of hydrogen peroxide is:



Balance the reaction. How many moles of water and oxygen are produced by the decomposition of 68.0 grams of hydrogen peroxide? How many molecules of water and oxygen are produced? How many grams of each product are formed?

$$68 \text{ g H}_2\text{O}_2 \times \frac{1 \text{ mol H}_2\text{O}_2}{34.01 \text{ g H}_2\text{O}_2} \times \frac{2 \text{ mol H}_2\text{O}}{2 \text{ mol H}_2\text{O}_2} = \boxed{2.00 \text{ mol H}_2\text{O}}$$

$$68 \text{ g H}_2\text{O}_2 \times \frac{1 \text{ mol H}_2\text{O}_2}{34.01 \text{ g H}_2\text{O}_2} \times \frac{1 \text{ mol O}_2}{2 \text{ mol H}_2\text{O}_2} = \boxed{1.00 \text{ mol O}_2}$$

$$68 \text{ g H}_2\text{O}_2 \times \frac{1 \text{ mol H}_2\text{O}_2}{34.01 \text{ g H}_2\text{O}_2} \times \frac{2 \text{ mol H}_2\text{O}}{2 \text{ mol H}_2\text{O}_2} \times \frac{1.0 \times 10^{24} \text{ molecules H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = \boxed{1.2 \times 10^{24} \text{ molecules H}_2\text{O}}$$

$$68 \text{ g H}_2\text{O}_2 \times \frac{1 \text{ mol H}_2\text{O}_2}{34.01 \text{ g H}_2\text{O}_2} \times \frac{1 \text{ mol O}_2}{2 \text{ mol H}_2\text{O}_2} \times \frac{1.0 \times 10^{24} \text{ molecules O}_2}{1 \text{ mol O}_2} = \boxed{6.0 \times 10^{23} \text{ molecules O}_2}$$

$$68 \text{ g H}_2\text{O}_2 \times \frac{1 \text{ mol H}_2\text{O}_2}{34.01 \text{ g H}_2\text{O}_2} \times \frac{2 \text{ mol H}_2\text{O}}{2 \text{ mol H}_2\text{O}_2} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = \boxed{36.02 \text{ g H}_2\text{O}}$$

$$68 \text{ g H}_2\text{O}_2 \times \frac{1 \text{ mol H}_2\text{O}_2}{34.01 \text{ g H}_2\text{O}_2} \times \frac{1 \text{ mol O}_2}{2 \text{ mol H}_2\text{O}_2} \times \frac{32.00 \text{ g O}_2}{1 \text{ mol O}_2} = \boxed{63.96 \text{ g O}_2}$$

3. If 156.0 grams of potassium metal reacts with excess water, potassium hydroxide and hydrogen gas are formed? What volume of hydrogen gas, in liters,

Be sure to balance the reaction first.



$$156 \text{ g K} \times \frac{1 \text{ mol K}}{39.09 \text{ g K}} \times \frac{1 \text{ mol H}_2}{2 \text{ mol K}} \times \frac{22.4 \text{ L H}_2}{1 \text{ mol H}_2} = \boxed{44.70 \text{ L H}_2}$$

$$156 \text{ g K} \times \frac{1 \text{ mol K}}{39.09 \text{ g K}} \times \frac{2 \text{ mol KOH}}{2 \text{ mol K}} \times \frac{56.11 \text{ g KOH}}{1 \text{ mol KOH}} = \boxed{87.37 \text{ g KOH}}$$

4. Determine the number of moles of carbon dioxide gas, water, and sodium chloride formed by the reaction of 42.0 grams of sodium bicarbonate (baking soda) reacting with excess hydrochloric acid. The reaction is:



$$42 \text{ g NaHCO}_3 \times \frac{1 \text{ mol NaHCO}_3}{84.01 \text{ g NaHCO}_3} \times \frac{1 \text{ mol CO}_2}{1 \text{ mol NaHCO}_3} = \boxed{0.5 \text{ mol CO}_2}$$

$$42 \text{ g NaHCO}_3 \times \frac{1 \text{ mol NaHCO}_3}{84.01 \text{ g NaHCO}_3} \times \frac{1 \text{ mol H}_2\text{O}}{1 \text{ mol NaHCO}_3} = \boxed{0.5 \text{ mol H}_2\text{O}}$$

$$42 \text{ g NaHCO}_3 \times \frac{1 \text{ mol NaHCO}_3}{84.01 \text{ g NaHCO}_3} \times \frac{1 \text{ mol NaCl}}{1 \text{ mol NaHCO}_3} = \boxed{0.5 \text{ mol NaCl}}$$

# Gas Stoichiometry Worksheet Answer Key With Work

**NARAYAN CHANGDER**



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## **Table of Contents Gas Stoichiometry Worksheet Answer Key With Work**

1. Understanding the eBook Gas Stoichiometry Worksheet Answer Key With Work
  - The Rise of Digital Reading Gas Stoichiometry Worksheet Answer Key With Work
  - Advantages of eBooks Over Traditional Books
2. Identifying Gas Stoichiometry Worksheet Answer Key With Work
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Gas Stoichiometry Worksheet Answer Key With Work
  - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Stoichiometry Worksheet Answer Key With Work
  - Personalized Recommendations
  - Gas Stoichiometry Worksheet Answer Key With Work User Reviews and Ratings
  - Gas Stoichiometry Worksheet Answer Key With Work and Bestseller Lists
5. Accessing Gas Stoichiometry Worksheet Answer Key With Work Free and Paid eBooks
  - Gas Stoichiometry Worksheet Answer Key With Work Public Domain eBooks
  - Gas Stoichiometry Worksheet Answer Key With Work eBook Subscription Services
  - Gas Stoichiometry Worksheet Answer Key With Work Budget-Friendly Options
6. Navigating Gas Stoichiometry Worksheet Answer Key With Work eBook Formats
  - ePub, PDF, MOBI, and More
  - Gas Stoichiometry Worksheet Answer Key With Work Compatibility with Devices
  - Gas Stoichiometry Worksheet Answer Key With Work Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Gas Stoichiometry Worksheet Answer Key With Work
  - Highlighting and Note-Taking Gas Stoichiometry Worksheet Answer Key With Work
  - Interactive Elements Gas Stoichiometry Worksheet Answer Key With Work
8. Staying Engaged with Gas Stoichiometry Worksheet Answer Key With Work

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Gas Stoichiometry Worksheet Answer Key With Work
- 9. Balancing eBooks and Physical Books Gas Stoichiometry Worksheet Answer Key With Work
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Gas Stoichiometry Worksheet Answer Key With Work
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Gas Stoichiometry Worksheet Answer Key With Work
  - Setting Reading Goals Gas Stoichiometry Worksheet Answer Key With Work
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Gas Stoichiometry Worksheet Answer Key With Work
  - Fact-Checking eBook Content of Gas Stoichiometry Worksheet Answer Key With Work
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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