

7. If you have 17 liters of gas at a temperature of 67 °C and a pressure of 88.89 atm, what will be the pressure of the gas if you raise the temperature to 94 °C and decrease the volume to 12 liters?

$$P_2 = \frac{P_1 V_1 T_2}{V_2 T_1} = \frac{88.89 \text{ atm} (17 \text{ L}) 367 \text{ K}}{(12 \text{ L}) 340 \text{ K}}$$

$$P_2 = \quad \text{atm}$$

8. I have an unknown volume of gas at a pressure of 0.5 atm and a temperature of 325 K. If I raise the pressure to 1.2 atm, decrease the temperature to 320 K, and measure the final volume to be 48 liters, what was the initial volume of the gas?

$$V_1 = \frac{P_2 V_2 T_1}{T_2 P_1} = \frac{1.2 \text{ atm} (48 \text{ L}) 325 \text{ K}}{320 \text{ K} (0.5 \text{ atm})}$$

$$V_1 = \quad \text{L}$$

9. If you have 21 liters of gas held at a pressure of 78 atm and a temperature of 900 K, what will be the volume of gas if you decrease the pressure to 45 atm and decrease the temperature to 750 K?

$$V_2 = \frac{78 \text{ atm} (21 \text{ L}) 750 \text{ K}}{900 \text{ K} (45 \text{ atm})}$$

$$V_2 = \quad \text{L}$$

10. If I have 2.9 L of gas at a pressure of 5 atm and a temperature of 50 °C, what will the temperature of the gas be if I decrease the volume of the gas to 2.4 L and decrease the pressure to 3 atm?

$$1/T_2 = \frac{P_1 V_1}{T_1 P_2 V_2} = \frac{(2.9 \text{ L}) 5 \text{ atm}}{323 \text{ K} (3 \text{ atm}) 2.4 \text{ L}}$$

$$T_2 = 160 \text{ K}$$

11. You have an unknown volume of gas held at a temperature of 115 K in a container with a pressure of 60 atm. If by increasing the temperature to 225 K and decreasing the pressure to 30 atm causes the volume of the gas to be 29 L, how many liters of gas did you start with?

$$V_1 = \frac{P_2 V_2 T_1}{T_2 P_1} = \frac{115 \text{ K} (29 \text{ L}) 30 \text{ atm}}{(225 \text{ K}) 60 \text{ atm}}$$

$$V_1 = \quad \text{L}$$

# Gas Laws Practice Calculations Key

**Michael Horsley, Sherwin Keith**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author names.

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**Thermodynamics Problem Solving in Physical Chemistry** Kathleen E. Murphy, 2020-03-23 Thermodynamics Problem Solving in Physical Chemistry Study Guide and Map is an innovative and unique workbook that guides physical chemistry students through the decision making process to assess a problem situation, create appropriate solutions, and gain confidence through practice solving physical chemistry problems. The workbook includes six major sections with 20-30 solved problems in each section that span from easy single objective questions to difficult multistep analysis problems. Each section of the workbook contains key points that highlight major features of the topic to remind students of what they need to apply to solve problems in the topic area. Key Features: Includes a visual map that shows how all the equations used in thermodynamics are connected and how they are derived from the three major energy laws. Acts as a guide in deriving the correct solution to a problem. Illustrates the questions students should ask themselves about the critical features of the concepts to solve problems in physical chemistry. Can be used as a stand alone product for review of Thermodynamics questions for major tests.

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## **Table of Contents Gas Laws Practice Calculations Key**

1. Understanding the eBook Gas Laws Practice Calculations Key
  - The Rise of Digital Reading Gas Laws Practice Calculations Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Gas Laws Practice Calculations Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in a Gas Laws Practice Calculations Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Laws Practice Calculations Key
  - Personalized Recommendations
  - Gas Laws Practice Calculations Key User Reviews and Ratings

- Gas Laws Practice Calculations Key and Bestseller Lists
- 5. Accessing Gas Laws Practice Calculations Key Free and Paid eBooks
  - Gas Laws Practice Calculations Key Public Domain eBooks
  - Gas Laws Practice Calculations Key eBook Subscription Services
  - Gas Laws Practice Calculations Key Budget-Friendly Options
- 6. Navigating Gas Laws Practice Calculations Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Gas Laws Practice Calculations Key Compatibility with Devices
  - Gas Laws Practice Calculations Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Gas Laws Practice Calculations Key
  - Highlighting and Note-Taking Gas Laws Practice Calculations Key
  - Interactive Elements Gas Laws Practice Calculations Key
- 8. Staying Engaged with Gas Laws Practice Calculations Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Gas Laws Practice Calculations Key
- 9. Balancing eBooks and Physical Books Gas Laws Practice Calculations Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Gas Laws Practice Calculations Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Gas Laws Practice Calculations Key
  - Setting Reading Goals Gas Laws Practice Calculations Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Gas Laws Practice Calculations Key
  - Fact-Checking eBook Content of Gas Laws Practice Calculations Key
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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