

## Mixed Gas Law Worksheet

Answers available at end

Name: \_\_\_\_\_ (show your work) period: \_\_\_\_\_

Convert:

$$100 \text{ mmHg} = \text{_____ kPa}$$

$$300 \text{ mmHg} = \text{_____ atm}$$

$$120 \text{ atm} = \text{_____ mmHg}$$

$$300 \text{ torr} = \text{_____ kPa}$$

$$25 \text{ kPa} = \text{_____ mmHg}$$

$$25 \text{ mmHg} = \text{_____ atm}$$

$$300^{\circ}\text{C} = \text{_____ K}$$

$$150\text{K} = \text{_____}^{\circ}\text{C}$$

$$5000^{\circ}\text{C} = \text{_____ K}$$

$$300\text{K} = \text{_____}^{\circ}\text{C}$$

$$-255^{\circ}\text{C} = \text{_____ K}$$

$$3500\text{K} = \text{_____}^{\circ}\text{C}$$

$$200^{\circ}\text{C} = \text{_____}^{\circ}\text{F}$$

$$95^{\circ}\text{F} = \text{_____}^{\circ}\text{C} = \text{_____ K}$$

$$45^{\circ}\text{F} = \text{_____}^{\circ}\text{C} = \text{_____ K}$$

- 1) A 70. liter sample of gas initially at  $280^{\circ}\text{C}$  is allowed to cool at constant pressure, what will the new volume be at  $0^{\circ}\text{C}$ ?

$$P_1 = \text{_____} \quad P_2 = \text{_____}$$

$$V_1 = \text{_____} \quad V_2 = \text{_____}$$

$$T_1 = \text{_____} \quad T_2 = \text{_____}$$

- 2) A 4.0 liter sample of gas initially at  $25^{\circ}\text{C}$  is allowed to cool at constant pressure, what will the new volume be at  $-5^{\circ}\text{C}$ ?

$$P_1 = \text{_____} \quad P_2 = \text{_____}$$

$$V_1 = \text{_____} \quad V_2 = \text{_____}$$

$$T_1 = \text{_____} \quad T_2 = \text{_____}$$

- 3) A buoyancy vest (used in scuba diving) is filled to a volume of 2.3 liters at a pressure of 1 atmosphere (at the surface where the temperature is a balmy 300K), what will the volume of air be in the vest when the diver dives to a depth of about 90 feet where the temperature is  $17^{\circ}\text{C}$  and the pressure is 4 atmospheres?

$$P_1 = \text{_____} \quad P_2 = \text{_____}$$

# Gas Laws And Scuba Diving Worksheet Answers

**Albert A Gayle**



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## **Table of Contents Gas Laws And Scuba Diving Worksheet Answers**

1. Understanding the eBook Gas Laws And Scuba Diving Worksheet Answers
  - The Rise of Digital Reading Gas Laws And Scuba Diving Worksheet Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Gas Laws And Scuba Diving Worksheet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Gas Laws And Scuba Diving Worksheet Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Laws And Scuba Diving Worksheet Answers

- Personalized Recommendations
  - Gas Laws And Scuba Diving Worksheet Answers User Reviews and Ratings
  - Gas Laws And Scuba Diving Worksheet Answers and Bestseller Lists
5. Accessing Gas Laws And Scuba Diving Worksheet Answers Free and Paid eBooks
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    - Gas Laws And Scuba Diving Worksheet Answers eBook Subscription Services
    - Gas Laws And Scuba Diving Worksheet Answers Budget-Friendly Options
  6. Navigating Gas Laws And Scuba Diving Worksheet Answers eBook Formats
    - ePub, PDF, MOBI, and More
    - Gas Laws And Scuba Diving Worksheet Answers Compatibility with Devices
    - Gas Laws And Scuba Diving Worksheet Answers Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Gas Laws And Scuba Diving Worksheet Answers
    - Highlighting and Note-Taking Gas Laws And Scuba Diving Worksheet Answers
    - Interactive Elements Gas Laws And Scuba Diving Worksheet Answers
  8. Staying Engaged with Gas Laws And Scuba Diving Worksheet Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Gas Laws And Scuba Diving Worksheet Answers
  9. Balancing eBooks and Physical Books Gas Laws And Scuba Diving Worksheet Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Gas Laws And Scuba Diving Worksheet Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Gas Laws And Scuba Diving Worksheet Answers
    - Setting Reading Goals Gas Laws And Scuba Diving Worksheet Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Gas Laws And Scuba Diving Worksheet Answers

- Fact-Checking eBook Content of Gas Laws And Scuba Diving Worksheet Answers
- Distinguishing Credible Sources

### 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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