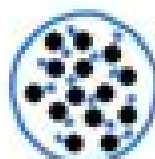




- This is colder because it is more condensed and so there is less movement.



- |                     |                                     |                                     |                 |
|---------------------|-------------------------------------|-------------------------------------|-----------------|
| Robert Boyle        | Jacques Charles                     | Guillaume Amontons                  | Amedeo Avogadro |
| $P_1 V_1 = P_2 V_2$ | $\frac{V_1}{T_1} = \frac{V_2}{T_2}$ | $\frac{P_1}{T_1} = \frac{P_2}{T_2}$ | $V \propto n$   |

Chemists combine all of the relationships seen in Models 1 and 2 into one law—the **Ideal Gas Law**. It is one equation that describes gas behavior and the relationship among all four variables,  $P$ ,  $V$ ,  $T$ , and  $n$ . In the Ideal Gas Law the proportionality constant is represented by the letter  $R$  (rather than the generic  $k$ ).

- $$P = RTnV$$

PT - RnV

$$PV = nRT$$

PTV = Rm

# Gas Variable Answers

**BM King**



## Gas Variable Answers:

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