

100) Carbon tetrachloride consists of one carbon and four chlorine atoms. How many chlorine atoms are present in 1.00 mole of CCl_4 molecules?

$$\text{one mole of } \text{CCl}_4 \text{ contains } 4 \text{ moles of } \text{Cl} \text{ atoms}$$

$$1 \text{ mole } \text{CCl}_4 \times \frac{4 \text{ moles Cl}}{1 \text{ mole } \text{CCl}_4} = 4 \text{ moles Cl}$$

$$\text{answer} = 4 \text{ moles Cl} \quad \text{100}$$

101) There are 4 monatomic gases in each buffer gas used for Phosor.

102) Describe how each of these gases is used.

Helium - for cooling the sample
 Argon - for shielding the sample
 Neon - for the sample
 Xenon - for the sample

103) Describe how each of these gases is used in the Phosor.

Helium - for cooling the sample
 Argon - for shielding the sample
 Neon - for the sample
 Xenon - for the sample

104) The chemical formula for the compound is $\text{Na}_2\text{S}_2\text{O}_3$.

105) The chemical formula for the compound is $\text{Na}_2\text{S}_2\text{O}_3$.

106) The chemical formula for the compound is $\text{Na}_2\text{S}_2\text{O}_3$.



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Tobias Bleicker



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