

Shop Manual

D68E-1 D68P-1

BULLDOZER

SERIAL NUMBERS	D68E-1	1001 and up
	D68E-1A (BRAZILIAN)	B1001 and up
	D68P-1	45590 and up
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$$\begin{array}{l} 1 \text{ N}_2 + 3 \text{ H}_2 \rightarrow 2 \text{ NH}_3 \\ 3 \text{ KClO}_3 \rightarrow 3 \text{ KCl} + 3 \text{ O}_2 \\ 2 \text{ NaCl} + \text{F}_2 \rightarrow 2 \text{ NaF} \\ 1 \text{ Cl}_2 + 4 \text{ H}_2 \text{O} \rightarrow 2 \text{ H}_2\text{O}_2 + 5 \text{ H}_2 \\ 1 \text{ Pb(OH)}_2 + 2 \text{ HCl} \rightarrow 2 \text{ H}_2\text{O} + 1 \text{ PbCl}_2 \\ 2 \text{ AlBr}_3 + 3 \text{ K}_2\text{SO}_4 \rightarrow 6 \text{ KBr} + 1 \text{ Al}_2\text{SO}_4 \\ 3 \text{ CH}_4 + 2 \text{ O}_2 \rightarrow 1 \text{ CO}_2 + 2 \text{ H}_2\text{O} \\ 8 \text{ C}_3\text{H}_8 + 5 \text{ O}_2 \rightarrow 3 \text{ CO}_2 + 4 \end{array}$$

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web 1 78 348 stoichiometry helps us determine how much substance is needed or is present things that can be measured are reactants and products mass molecular weight chemical equations formulas stoichiometric coefficient the stoichiometric coefficient or stoichiometric number is the number of molecules that participate in the reaction

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$$\begin{array}{l} 1 \text{ N}_2 + \text{F}_2 \rightarrow 2 \text{ NF}_3 \\ 2 \text{ C}_6\text{H}_{10} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O} \\ 3 \text{ HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2 \\ \text{GaBr}_3 + \text{NaBr} \rightarrow \text{SNO}_4 + \text{NaSO}_3 + \text{Ga}_2\text{SO}_3 + 5 \text{ NF}_3 + \text{SnF}_2 + \text{N}_2\text{O}_3 \end{array}$$
 using the equation from problem 2 above answer the following

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$$\begin{array}{l} 1 \text{ N}_2 + 3 \text{ F}_2 \rightarrow 2 \text{ NF}_3 \\ 2 \text{ C}_6\text{H}_{10} + 17 \text{ O}_2 \rightarrow 12 \text{ CO}_2 + 10 \text{ H}_2\text{O} \\ 3 \text{ HBr} + 1 \text{ KHCO}_3 \rightarrow 1 \text{ H}_2\text{O} + 1 \text{ KBr} + \text{CO}_2 \\ 2 \text{ GaBr}_3 + 3 \text{ Na}_2\text{SO}_3 \rightarrow 1 \text{ Ga}_2\text{SO}_3 + 3 \text{ SO}_3 + 6 \text{ NaBr} + 3 \text{ SNO}_2 + 3 \text{ NF}_3 + \text{SnF}_2 + 1 \text{ N}_2\text{O}_3 \end{array}$$
 using the following equation $2 \text{ NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{ H}_2\text{O} + \text{Na}_2\text{SO}_4$

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assuming CeCl_2 CuCl_2 CuCl_2 is available in excess

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 $\text{Br}_2 + \text{C}_6\text{H}_{10} \rightarrow \text{C}_6\text{H}_8 + \text{H}_2\text{O}$
 $\text{H}_2\text{O} + \text{KBr} + \text{CO}_2 \rightarrow \text{K}_2\text{SO}_3 + \text{Ga}_2\text{SO}_3 + \text{NaBr}$
 using the following equation
 $\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow \text{H}_2\text{O} + \text{Na}_2\text{SO}_4$

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 $\text{N}_2 + \text{F}_2 \rightarrow \text{NF}_3$
 $\text{CaH}_2 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O} + \text{HBr} + \text{KHCO}_3$
 $\text{H}_2\text{O} + \text{KBr} + \text{CO}_2 \rightarrow \text{Ga}_2\text{SO}_3 + \text{NaBr} + \text{SNO} + \text{NF}_3$
 SNO_2

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web may 28 2020 write the balanced equation for the reaction and calculate how much SiO_2 is required to produce 3.00 kg of SiC
 $\text{SiO}_2 + 3\text{C} \rightarrow \text{SiC} + 2\text{CO}$

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 $\text{N}_2 + \text{F}_2 \rightarrow \text{NF}_3$
 $\text{C}_6\text{H}_{10} + \text{O}_2 \rightarrow \text{C}_6\text{H}_8 + \text{H}_2\text{O}$
 $\text{KBr} + \text{CO}_2 \rightarrow \text{Ga}_2\text{SO}_3 + \text{NaBr} + \text{SNO} + \text{NF}_3$
 solve the following stoichiometry grams problems 6

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