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Fig. 3.1

(a) The symbol equation for the reaction can be represented as shown in Fig. 3.1.

Fig. 3.1

Table 3.1 shows some bond energies.

bond	value	unit
bond energy in kJ/mol	440	kJ/mol

Use the bond energies in Table 3.1 and ΔH to calculate the bond energy of one C-H bond in kJ/mol.

Use the following steps.

- Calculate the energy needed to break bonds in the reactants.

$$(500) + 3(440)$$

$$440 + 3(440)$$

$$1660 \text{ kJ/mol}$$
- Calculate the energy released when bonds form in the products.

$$\text{bond broken} - \text{bond formed} = -90$$

$$1660 - x = -90$$

$$x = 1660 + 90$$

$$1750 \text{ kJ/mol}$$
- Calculate the energy of one C-H bond.

$$6 \times \text{C-H} = 1750$$

$$\text{C-H} = \frac{1750}{6}$$

$$292 \text{ kJ/mol}$$

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Léo Pomar

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