3.

CHAPTER 9 REVIEW

Stoichiometry

SECTION 2: Ideal Stoichiometric Calculations

Please make sure to use correct formulas for the calculations.

The following equation represents a laboratory preparation
for oxygen gas:

 $2KClO_2(s) \rightarrow 2KCl(s) + 3O_2(g)$

How many moles of O₂ form if 3.0 mol of KClO₃ are totally consumed?

Show full workings of your answer in this box

Given the following equation:

 $H_2(g) + F_2(g) \rightarrow 2HF(g)$ How many grams of HF gas are produced as 5 mol of fluorine react?

Show full workings of your answer in this box

Water can be made to decompose into its elements by using electricity according to the following equation:

 $2H_2O(l) \rightarrow 2H_2(g) + O_2(g)$

How many grams of O₂ are produced when 0.033 mol of water decompose?

Show full workings of your answer in this box

Sodium metal reacts with water to produce NaOH according to the following equation:

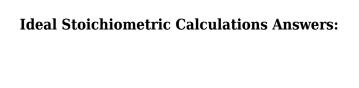
 $2Na(s) + 2H_2O(I) \rightarrow 2NaOH(aq) + H_2(g)$

How many grams of NaOH are produced if 20.0 g of sodium metal react with excess oxygen?

Show full workings of your answer in this box

Ideal Stoichiometric Calculations Answers

Rachel S Tattersall



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