

## Conclusion

Through gravimetric analysis, we were able to determine the percentage of chloride in the dried sample. One of the most important sources of error in this experiment is the nature of the precipitate. The result of adding the precipitate in a hasty manner instead of doing it slowly will be ended up form a colloidal precipitate which was the reason why solution in the beaker does not readily distinguish the separation of the solvent to precipitate. Besides that, Using high concentration of substance that was measured and weight and precipitant can also lead to form of not good precipitate or it take longer than usual to get a good result. To achieve good precipitate, one must constantly stir the solution while heating because it is important to digest the colloidal sample. Other than that, the sample may absorb moisture in the air when it was cooling down may lead to an increasing weight compared to the original. It is important to note about systematic error when carry out the experiment. Nevertheless, it proved that gravimetric analysis can accurately determine the result of the experiment as much as stoichiometric method.

## Recommendation

- Make sure the experiment must be carry out in one go.
- Keep stirring the solution while heating to avoid coagulation.
- The crucible must be completely dried before weighting it to avoid measuring error.

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## Appendices

1. Equation:  $\text{Na}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \longrightarrow \text{NaCl}(\text{aq})$

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**DP Hallahan**



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