

## Half-Life

If 100.0 g of carbon-14 decays until only 25.0 g of carbon is left after 11,460 y, what is the half-life of carbon-14?

1. List the given and unknown values.

*Given:* initial mass of sample = 100.0 g  
 final mass of sample = 25.0 g  
 total time of decay = 11 460 y

*Unknown:* number of half-lives = ? half-lives  
 half-life = ? y

2. Write down the equation relating half-life, the number of half-lives, and the decay time, and rearrange it to solve for half-life.

$$\text{total time of decay} = \text{number of half-lives} \times \frac{\text{number of years}}{\text{half-life}}$$

$$\frac{\text{number of years}}{\text{half-life}} = \frac{\text{total time of decay}}{\text{number of half-lives}}$$

3. Calculate how many half-lives have passed during the decay of the 100.0 g sample.

$$\text{fraction of sample remaining} = \frac{\text{final mass of sample}}{\text{initial mass of sample}} = \frac{25.0 \text{ g}}{100.0 \text{ g}} = \frac{1}{4}$$

$$\text{after one half-life} = \frac{1}{2}; \text{ after two half-lives} = \frac{1}{2} \times \frac{1}{2} = \frac{1}{4} \text{ of sample}$$

Two half-lives have passed.

4. Solve for the half-life.

$$\frac{\text{number of years}}{\text{half-life}} = \frac{11\,460 \text{ y}}{2 \text{ half-lives}} = \frac{5730 \text{ y}}{\text{half-life}}$$

$$\text{half-life of carbon-14} = 5730 \text{ y}$$

## Practice

- What is the half-life of a 100.0 g sample of nitrogen-16 that decays to 12.5 g of nitrogen-16 in 21.6 s?
- All isotopes of technetium are radioactive, but they have widely varying half-lives. If an 800.0 g sample of technetium-99 decays to 100.0 g of technetium-99 in 639 000 y, what is its half-life?
- A 208 g sample of sodium-24 decays to 13.0 g of sodium-24 within 60.0 h. What is the half-life of this radioactive isotope?

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**J Dewey**



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### **Table of Contents Math Skills Half Life Worksheet Answers**

1. Understanding the eBook Math Skills Half Life Worksheet Answers
  - The Rise of Digital Reading Math Skills Half Life Worksheet Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Math Skills Half Life Worksheet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Math Skills Half Life Worksheet Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Math Skills Half Life Worksheet Answers
  - Personalized Recommendations
  - Math Skills Half Life Worksheet Answers User Reviews and Ratings
  - Math Skills Half Life Worksheet Answers and Bestseller Lists

5. Accessing Math Skills Half Life Worksheet Answers Free and Paid eBooks
  - Math Skills Half Life Worksheet Answers Public Domain eBooks
  - Math Skills Half Life Worksheet Answers eBook Subscription Services
  - Math Skills Half Life Worksheet Answers Budget-Friendly Options
6. Navigating Math Skills Half Life Worksheet Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Math Skills Half Life Worksheet Answers Compatibility with Devices
  - Math Skills Half Life Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Math Skills Half Life Worksheet Answers
  - Highlighting and Note-Taking Math Skills Half Life Worksheet Answers
  - Interactive Elements Math Skills Half Life Worksheet Answers
8. Staying Engaged with Math Skills Half Life Worksheet Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Math Skills Half Life Worksheet Answers
9. Balancing eBooks and Physical Books Math Skills Half Life Worksheet Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Math Skills Half Life Worksheet Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Math Skills Half Life Worksheet Answers
  - Setting Reading Goals Math Skills Half Life Worksheet Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Math Skills Half Life Worksheet Answers
  - Fact-Checking eBook Content of Math Skills Half Life Worksheet Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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