

ICSE Board Chemistry Syllabus for Class 9

COURSE STRUCTURE CLASS IX

SUBJECT-CHEMISTRY

1. Matter and Its Composition: Law of Conservation of mass

(i) Explanation of change of state of matter on the basis of Kinetic Theory of Matter. Main postulates of Kinetic Theory of Matter and explanation of change of state on the basis of. Inter-particle space and interparticle attraction and collision.

(ii) Law of Conservation of Mass Statement and explanation with examples.

2. Study of Gas Laws

(i) The behaviour of gases under changes of temperature and pressure; explanation in terms of molecular motion (particles, atoms, molecules); Boyle's Law and Charles' Law; absolute zero; gas equation; simple relevant calculations.

The behaviour of gases under changes of temperature and pressure; explanation in terms of molecular motion (particles, atoms, molecules). Boyle's Law (statement, mathematical form, simple calculations). Charles' Law (statement, mathematical form, simple calculations). Absolute zero; Kelvin scale of temperature. Gas equation $P_1 V_1 / T_1 = P_2 V_2 / T_2$; simple relevant calculations based on gas equation.

(ii) Relationship between Kelvin Scale and Celsius Scale of temperature; Standard temperature and pressure. Conversion of temperature from Celsius Scale to Kelvin scale and vice versa. Standard temperature and pressure. (simple calculations).

3. Elements, Compounds and Mixtures

(i) General characteristics and differences between elements, compounds and mixtures.

Reasons for considering a substance as an element, compound or mixture may be given to make the concepts clear.

(ii) Types of mixtures: of two solids, a solid and a liquid, two liquids, liquid and gas, two gases.

Definition of mixture; each type of mixture should be shown to the students (including both homogeneous and heterogeneous types) - true solution, suspension and colloidal solution to make the concepts clear.

(iii) Separation of mixtures involving - use of a solvent, filtration, evaporation and distillation, fractional distillation, simple paper chromatography Centrifugation immiscible liquid.

The following examples should be used to illustrate the principles of separation of mixtures by using following methods.

(a) use of solvent and filtration (e.g. sodium chloride + sand, water as solvent), carbon and sulphur (Carbon tetra chloride as solvent)

(b) evaporation e.g. sodium chloride from its aqueous solution.

(c) distillation e.g. purification of water containing dissolved solids.

(d) fractional distillation involves the difference in boiling points of liquids e.g. benzene + toluene.

(e) simple paper chromatography (limited to separation of colouring matter in ink);

(f) Centrifugation (involving separation of cream from milk).

(g) immiscible liquids (separating funnel e.g. water + carbon tetra chloride).

4. The language of Chemistry

Symbol of an element; valency; formulae of radicals and formulae of compounds. Balancing of simple chemical equations.

Symbol - definition; symbols of the elements used often. Valency - definition; hydrogen combination and number of valence electrons of the metals and non-metals; mono, di, tri and tetra valent elements. Radicals - definition of radicals; formula and valencies of the radicals and formula of compounds.

Chemical equation - definition and examples of chemical equations with one reactant and two or three products, two reactants and one product, two reactants and two products and two reactants and three or four products; balancing of equations. (By partial equation method and hit and trial method)

5. Physical and Chemical Changes

(i) Definitions and distinction between Physical and Chemical changes. Simple experiments like dissolution of sugar

in water, burning of paper should be shown to make the concepts of physical and chemical change clear. More examples of such type may be given.

(ii) Conditions for chemical change. Close contact, heat, light, electricity, pressure, catalysts with examples.

(iii) Types of chemical change. Direct combination; decomposition; displacement; double decomposition with examples.

(iv) Energy changes in a chemical change. Exothermic and endothermic reactions with examples - evolution/absorption of heat, light and electricity.

(v) Burning: Definition and conditions of burning. Definition; conditions for burning (combustible substance, supporter of combustion and ignition temperature); comparison of respiration and burning; burning of magnesium or candle to show that substances gain weight on burning; students to be made aware of how the balance of O_2 and CO_2 is maintained in nature.

6. Water

Water as a compound and as a universal solvent; its physical and chemical properties.

Why is water considered a compound? Chief physical properties should include: density, b.p., m.p. Experiment to show that the water we drink, contains dissolved solids and dissolved gases (air); their significance. Solutions as 'mixtures' of solids in water; saturated solutions; qualitative

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R Sandford



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