



Hp 1500 Printer Manual

KJ Lindholm-Leary



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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^- + \text{Au} \rightarrow 2\text{e}^- + 2\text{H}_2\text{O} + 2\text{H}_2\text{O}_2$ make electrons equal convert to basic solution $4\text{CN}^- + 2\text{Au} \rightarrow 2\text{e}^- + 2\text{H}_2\text{O} + 2\text{H}_2\text{O}_2$ multiplied by a factor of 2 $2\text{e}^- + 2\text{H}_2\text{O} + 2\text{H}_2\text{O}_2$ the final answer

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equations one for oxidation and one for reduction the equation is balanced by adjusting coefficients and adding H^+ and e^- in this order balance the atoms in the equation apart from O and H

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively $\text{C}_6\text{H}_8\text{O}_6 + 2\text{HNO}_2 \rightarrow \text{C}_6\text{H}_6\text{O}_6 + 2\text{H}_2\text{O} + 2\text{NO}$ choose 1 answer $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 a $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 b NO_2 and $\text{C}_6\text{H}_8\text{O}_6$ c $\text{C}_6\text{H}_8\text{O}_6$ and NO d NO and $\text{C}_6\text{H}_8\text{O}_6$